
Physical Chemistry

Third Edition

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University of Maryland



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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1/267 per degree in 1874 Victor Regnault obtained better experimental results showing this

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be
22 414 l 22 414 l we can substitute 101 325 kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k
for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of
gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the
standard molar volume figure 9 18

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used and with it we can explore it by seeing how the pressure and volumes changed in the tubes each time that we

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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101.36 kpa when the temperature was at 23.2 °C

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature °C temperature K constant k p t 95.82 kpa 16.0 °C 289 K 0.33 97.90 kpa 19.0 °C 292 K 0.34 98.97 kpa 99.22 kpa 101.4 kpa 19.1 °C

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