

CHEMISTRY: PERIODIC LAW "SCRAMBLE" LAB

WHAT TO TURN IN:

Unscrambled table

Questions #1-6

Introduction

The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869. The basis of his arrangement was the atomic masses of the elements. This approach proved incorrect as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the basis of the atomic numbers of the elements. In accordance with Moseley's revision, the periodic law states the properties of the elements are periodic functions of their atomic numbers.

Each of the known elements has its own set of properties. These range from solid to gas, lustrous to dull, low to high melting points, various colors, etc. The elements are arranged within the periodic table into groups or families (vertical columns) and periods or rows (horizontal rows). This arrangement reflects the repeating nature of the properties of the elements.

In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table.

Objectives

- To arrange the elements in Groups IA-VIIIA (1, 2, 13-18) according to a list of clues and your knowledge of periodic properties
- To explain the trends of properties in groups and periods

Equipment: scissors, glue stick, blank table, scrambled table of elements

Procedure

1. Obtain a "scrambled" table. Each block on the table represents a different element from Groups IA-VIIIA, (1-18)
2. Cut out blocks A-Z. Use the clues given at the end of the procedure and arrange the elements in their proper order the blank table provided by your teacher. When you are sure of a correct placement, glue the square in place. Elements A-Z will fill the first four rows.
3. Cut out the remaining 16 blocks (these have a * instead of a code letter). Use the information provided in each block and your knowledge of periodic properties to arrange these elements in their proper position on the blank table. Glue the blocks in place in rows five and six. **THESE WILL BE AN EXTRA OPTIONAL GRADE AND WILL NOT AFFECT YOUR LAB SCORE IF YOU GET ANY INCORRECT OR DO NOT DO THEM.**
4. Some information is missing from each block. Predict the values for the missing items from the location of the element on the periodic table. Place your predictions on the table. (You may use a periodic table *only* to determine the symbol for each element.) You must include the element symbol, atomic number, and oxidation number (charge):

GROUP NUMBER

Group IA (1)
Group IIA (2)
Group IIIA (13)
Group IVA (14)
Group VA (15)
Group VIA (16)
Group VIIA (17)
Group VIIIA/0 (18)

GENERAL CHARGE (OXIDATION NUMBER) OF ION

1+
2+
3+
varies (mixed)
3-
2-
1-
0 (none)

Periodic Law Scramble Lab Answers

Henry Bent



Periodic Law Scramble Lab Answers:

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range of variation less than one tenth of a unit The degree of probability that this is purely accidental is found to be only equal to 1 10^{97.8} This seems to illustrate the point that not only is Prout's law not as yet absolutely overturned but that a heavy and

The Periodic Law David Johnson, 1984 **The Developement of the Periodic Law** Francis Preston Venable, 1974 **Mendeleev on the Periodic Law** Dmitri Ivanovich Mendeleev, 2013-04-25 By the dawn of the nineteenth century elements had been defined as basic building blocks of nature resistant to decomposition by chemical means In 1869 the Russian chemist Dmitri Ivanovich Mendeleev organized the discord of the elements into the periodic table assigning each element to a row with each row corresponding to an elemental category The underlying order of matter hitherto only dimly perceived was suddenly clearly revealed This is the first English language collection of Mendeleev's most important writings on the periodic law Thirteen papers and essays divided into three groups reflect the period corresponding to the initial establishment of the periodic law three papers 1869-71 a period of priority disputes and experimental confirmations five papers 1871-86 and a final period of general acceptance for the law and increasing international recognition for Mendeleev five papers 1887-1905 A single easily accessible source for Mendeleev's principle papers this volume offers a history of the development of the periodic law written by the law's own founder Quiz Yourself Clever! The Periodic Table DK, 2024-10-03 Get to grips with the periodic table with this fact packed quiz book Learn all the ingredients that make up the universe in this quiz book for children aged 9 that breaks down the periodic table Quiz Yourself Clever The Periodic Table goes through all 118 elements in an accessible and kid friendly way showing the pure form of every element through an eye catching image surrounded by quiz style questions This periodic table quiz book for children offers Fast and fun learning through the use of detailed images and interesting facts Hundreds of quick fire questions for children to quiz themselves and challenge friends and family Information that has been endorsed and authenticated by experts Vivid pictures and fast facts will allow children to gather a wealth of knowledge in a fun way Turn the page to find the answers contained in a handy dataset from its atomic number and mass to its melting point and when it was discovered Images of how it appears in nature and how we use it give a fuller understanding of each element More in the series If you enjoyed Quiz Yourself Clever The Periodic Table then why not test yourself other quiz books to boost your knowledge like Quiz Yourself Clever Rocks Minerals to get to grips with the fascinating rock and mineral forms all over the world or Animals of the World to learn about the most fascinating wildlife on our planet On the Discovery of the Periodic Law, and on Relations Among the Atomic Weights John A. R. Newlands, 1884 **On the Discovery of the Periodic Law and on Relations Among the Atomic Weights** John A. R. Newlands, 1844 The Basics of the Periodic Table Leon Gray, 2013-12-15 Provides basic information on the periodic table Includes biographical information on Dmitri Mendeleev color photographs and diagrams sidebars a glossary and further reading sources *Energy Wave of the Periodic Table* William C. Scott, 2007 Definitive Periodic Law is revealed in arrangement of new Periodic Table repeating sequential numbers of protons discovered in Groups 1 through 18 in the

elements of the new ENERGY WAVE of the Periodic Table The elements in the new ENERGY WAVE of the Periodic Table are given in the ground state which is one electron for each proton This arrangement provided a unique opportunity to observe the nucleus of the elements By incorporating the sequential numbers of protons underlying the Energy Levels K L M N O P Q in shell blocks s p d and f of the Group elements it revealed what had been hidden and veiled in the complexity of electron configurations Sequential numbers of protons are observed to repeat in the Group elements from period to period This is the true revealed energy force creating the similar physical and chemical properties of Groups 1 through 18 from period to period in the Periodic Table The ENERGY WAVE of the Periodic Table had revealed Definitive Periodic Law Definitive Periodic Law is the number of protons underlying the Energy Levels K L M N O P Q in the nucleus of the Elements These sequential numbers of protons repeat in shell blocks s p d and f forming groups that have similar physical and chemical properties from period to period These sequential numbers of protons are the cornerstones of the nucleus and provide the atomic orbitals of the electrons the foundation for their spatial relationship to the nucleus as described by the azimuthal angular and magnetic numbers of quantum chemistry These sequential numbers of protons are very important as they reveal new explanation to chemical bond angles and the molecular geometry and structure of molecules **The Periodic Law**
Thomas Carnelley,1885 The periodic table Brian J. Knapp,1998 Twelve Volume Reference for beginning chemistry students

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