

## Periodic Trends Guided Note Sheet

Name \_\_\_\_\_

Period \_\_\_\_\_

Date \_\_\_\_\_

Fill in the blanks. Some blanks have word choices written after them to help you. Refer to the drawings and questions from the first page to help you with the fill in the other pages.

### I. Atomic Radius in a Group

- A. As you move down a group, the nucleus becomes \_\_\_\_\_ (stronger, weaker) because more \_\_\_\_\_ (electrons, protons) are being added to it. This stronger nucleus can pull the electrons in more tightly, making the atoms **smaller; smaller**.
- B. \_\_\_\_\_ (Protons, Electrons) are also added to the atom but are added to \_\_\_\_\_ (larger, smaller, the same) energy levels resulting in **larger** atoms.
- C. The increased number of inner electrons shield (or block) the \_\_\_\_\_ (positive, negative) pull of the nucleus.
- D. So, electron shielding \_\_\_\_\_ (increases, decreases) causing the electrons on the outermost energy levels to be even further from the nucleus than expected.
- E. This causes the atoms at the bottom of a Group to be even \_\_\_\_\_ (larger, smaller) than expected.
- F. Therefore, atomic radius \_\_\_\_\_ (increases, decreases) as you move down a group.

### II. Atomic Radius in a Period

- A. As you move right across a Period, the nucleus becomes \_\_\_\_\_ (stronger, weaker) because more \_\_\_\_\_ (electrons, protons) are being added to it.
- B. \_\_\_\_\_ (Protons, Electrons) are also added but they are added to \_\_\_\_\_ (the same, a larger, a smaller) energy level.
- C. This causes the amount of electron shielding from electrons on inner energy levels to remain \_\_\_\_\_ (constant, varied).
- D. Therefore, electrons are pulled in \_\_\_\_\_ (more, less) tightly because they feel the pull from nucleus more strongly.
- E. This causes the atoms on the right of the Periodic Table to be \_\_\_\_\_ (larger, smaller) than those on the left.
- F. So, atomic radius \_\_\_\_\_ (increases, decreases) as you move across a Period from left to right.

### III. Sizes of an Ion Versus Sizes of an Atom (Digital Textbook—Section 11.2)

- A. Metal cations are \_\_\_\_\_ (smaller, larger) than the neutral atoms.
- B. because \_\_\_\_\_
- C. Nonmetal anions are \_\_\_\_\_ (smaller, larger) than the neutral atoms.
- D. because \_\_\_\_\_

### IV. Ionization Energy in a Group

- A. Ionization energy is the energy required to \_\_\_\_\_ the outermost electron from an atom.
- B. As you move down a Group, the nucleus becomes \_\_\_\_\_ (stronger, weaker) because there are \_\_\_\_\_ (more, fewer) protons in it.
- C. **Constant**, as you move down a group, there are \_\_\_\_\_ (fewer, more) energy levels making the atom \_\_\_\_\_ (larger, smaller).
- D. In atoms at the top of a group, the outermost electrons are closer to the nucleus so the electrons can feel the pull of nucleus \_\_\_\_\_ (more, less) strongly.
- E. The outermost electrons also feel the pull of the nucleus less strongly because there is \_\_\_\_\_ (less, more) electron shielding from the \_\_\_\_\_ (decreased, increased) number of inner electrons.
- F. Therefore, \_\_\_\_\_ (more, less) energy is required to remove the outermost electron from a larger atom.

# Periodic Trends Review Notes Sheet Answers

**JG Myers**



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