

# The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

Groups

↓ ↓  
(Columns)

The elements in each group have the same number of valence electrons.

1 <sup>st</sup> Period → (1 shell) ☉	Group 1 (1 valence e <sup>-</sup> )	1 H Hydrogen	Group 2 (2 valence electrons)	4 Be Beryllium	Group 3 (3 valence electrons)	5 B Boron	Group 4 (4 valence electrons)	6 C Carbon	Group 5 (5 valence electrons)	7 N Nitrogen	Group 6 (6 valence electrons)	8 O Oxygen	Group 7 (7 valence electrons)	9 F Fluorine	Group 8 (8 valence electrons)	10 Ne Neon	Except Helium with 2 valence e <sup>-</sup>
2 <sup>nd</sup> Period → (2 shells) ☉☉		3 Li Lithium															
3 <sup>rd</sup> Period → (3 shells) ☉☉☉		11 Na Sodium		12 Mg Magnesium		13 Al Aluminum		14 Si Silicon		15 P Phosphorus		16 S Sulfur		17 Cl Chlorine		18 Ar Argon	
4 <sup>th</sup> Period → (4 shells) ☉☉☉☉		19 K Potassium		20 Ca Calcium		31 Ga Gallium		32 Ge Germanium		33 As Arsenic		34 Se Selenium		35 Br Bromine		36 Kr Krypton	

Periods →  
(rows)

The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

# Periodic Table Notes For Middle School

**Gary W Schneider**



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## **Table of Contents Periodic Table Notes For Middle School**

1. Understanding the eBook Periodic Table Notes For Middle School
  - The Rise of Digital Reading Periodic Table Notes For Middle School
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Notes For Middle School
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Notes For Middle School
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Notes For Middle School
  - Personalized Recommendations
  - Periodic Table Notes For Middle School User Reviews and Ratings
  - Periodic Table Notes For Middle School and Bestseller Lists
5. Accessing Periodic Table Notes For Middle School Free and Paid eBooks
  - Periodic Table Notes For Middle School Public Domain eBooks
  - Periodic Table Notes For Middle School eBook Subscription Services
  - Periodic Table Notes For Middle School Budget-Friendly Options
6. Navigating Periodic Table Notes For Middle School eBook Formats

- ePub, PDF, MOBI, and More
- Periodic Table Notes For Middle School Compatibility with Devices
- Periodic Table Notes For Middle School Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Table Notes For Middle School
  - Highlighting and Note-Taking Periodic Table Notes For Middle School
  - Interactive Elements Periodic Table Notes For Middle School
- 8. Staying Engaged with Periodic Table Notes For Middle School
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Notes For Middle School
- 9. Balancing eBooks and Physical Books Periodic Table Notes For Middle School
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Table Notes For Middle School
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Notes For Middle School
  - Setting Reading Goals Periodic Table Notes For Middle School
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Notes For Middle School
  - Fact-Checking eBook Content of Periodic Table Notes For Middle School
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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