

### Quantum Numbers

- \_\_\_\_\_
- Used to \_\_\_\_\_ an \_\_\_\_\_ in an \_\_\_\_\_

#### $n$

- \_\_\_\_\_
- Represents \_\_\_\_\_ energy level of \_\_\_\_\_
- \_\_\_\_\_ # of \_\_\_\_\_ in an \_\_\_\_\_

Example: What is the maximum number of electrons that can be in the \_\_\_\_\_ main energy level?

#### $l$

- The \_\_\_\_\_
- Describes the \_\_\_\_\_ within an \_\_\_\_\_
- \_\_\_\_\_ of orbital \_\_\_\_\_ possible in \_\_\_\_\_

### Orbital Shapes

designated \_\_\_\_\_

- level 1: \_\_\_\_\_
- level 2: \_\_\_\_\_
- level 3: \_\_\_\_\_
- level 4: \_\_\_\_\_

### How many electrons can each sublevel hold?

$s = 1 \text{ orbital} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$   
 $p = 3 \text{ orbitals} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$   
 $d = 5 \text{ orbitals} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$   
 $f = 7 \text{ orbitals} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$

# Note Taking Guide Episode 3answers Chemistry

**Clemens Wendtner**



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