

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
Mg(OH) ₂	<u>base</u>	H ₃ PO ₄	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO ₃	<u>salt</u>	Al(OH) ₃	<u>base</u>
HFO ₄	<u>acid</u>	KC ₂ H ₃ O ₂	<u>salt</u>
Ba(OH) ₂	<u>base</u>	NaCl	<u>salt</u>

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