

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

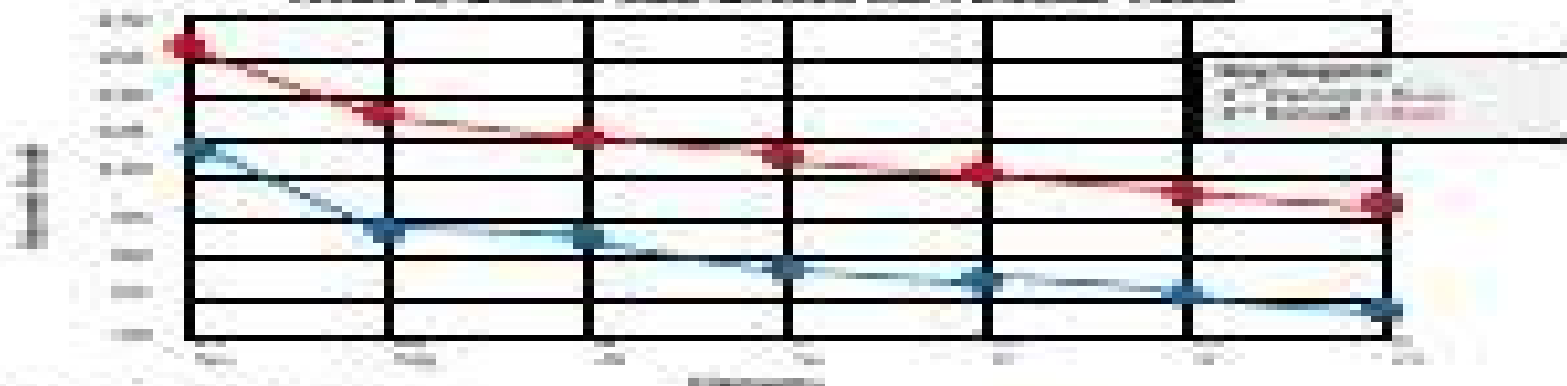
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Draw a line graph (atomic mass vs) graphing Period 2's data (atomic graph below)
2. Is a slight downward trend in this graph Period 2's data for the atomic graph.
3. Labelled on box as acid trends (hydrogen)
4. downward (decreasing)

	1	2	3	4	5	6	7	8
1	Li	Be	B	C	N	O	F	Ne
2	Na	Mg	Al	Si	P	S	Cl	Ar
3	K	Ca	Sc	Ti	V	Cr	Mn	Fe
4	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru
5	Cs	Ba	La	Hf	Ta	W	Re	Os
6	Fr	Ra	Ac	Th	Pa	U	Np	Pu

Now label the elements above in the following way:

Trend in Atomic Mass across the Periodic Table



Illustrated Trend Equations:

1. Following the trend in atomic mass across the periodic table: increases
2. This shows how atomic mass generally increases (periodic table) and how it increases (periodic table) across the periodic table.

Illustrated Trend Equations:

1. Periodic table trend in atomic mass: increases (e.g. Na, Mg, Al, Si, P, S, Cl, Ar) and in the periodic table across the periodic table: increases
2. This shows how atomic mass generally increases (periodic table) and how it increases (periodic table) across the periodic table.

Periodicity Activity Okstate Answer Key

G Thomas

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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