

# Step 8. Periodic table trends



Atomic radius

1																	18
1 H 1.008																	2 He 4.0026
3 Li 6.94	4 Be 9.0122											10 Ne 20.180					
11 Na 22.990	12 Mg 24.305									17 Cl 35.45	18 Ar 39.948						
19 K 39.098	20 Ca 40.078	21 Sc 44.956	22 Ti 47.867	23 V 50.942	24 Cr 51.996	25 Mn 54.938	26 Fe 55.845	27 Co 58.933	28 Ni 58.693	29 Cu 63.546	30 Zn 65.38	35 Br 79.904	36 Kr 83.798				
37 Rb 85.468	38 Sr 87.62	39 Y 88.906	40 Zr 91.224	41 Nb 92.906	42 Mo 95.94	43 Tc 98.906	44 Ru 101.07	45 Rh 102.905	46 Pd 106.42	47 Ag 107.868	48 Cd 112.411	51 Sb 121.757	52 Te 127.60	53 I 126.905	54 Xe 131.29		
55 Cs 132.905	56 Ba 137.33	57-71 Lanthanide series	72 Hf 178.49	73 Ta 180.948	74 W 183.84	75 Re 186.207	76 Os 190.23	77 Ir 192.22	78 Pt 195.084	79 Au 196.967	80 Hg 200.59	81 Tl 204.38	82 Pb 207.2	83 Bi 208.98	84 Po (209)	85 At (210)	86 Rn (222)
87 Fr (223)	88 Ra (226)	f-Block series										115 Sb (223)	116 Te (228)	117 Ts (289)	118 Og (286)		
		89 La 138.905	90 Ce 140.12	91 Pr 140.908	92 Nd 144.24	93 Pm (145)	94 Pu 239.04	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (260)	

Electronegativity

Ionization Energy

Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- Atomic radius of K is HIGHER than atomic radius of Br
- Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br

# Periodic Trends Answer Key Section Review

**J Ma**



**Periodic Trends Answer Key Section Review:**

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