

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 4 Volatile metals
- 5 Lanthanoids
- 6 Chalcogens
- 7 Halogens
- 8 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

1	2											13	14	15	16	17	18
1 H 1.01 Hydrogen	2 He 4.003 Helium											3 B 10.81 Boron	4 C 12.01 Carbon	5 N 14.01 Nitrogen	6 O 16.00 Oxygen	7 F 19.00 Fluorine	8 Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.012 Beryllium											9 Al 26.98 Aluminium	10 Si 28.09 Silicon	11 P 30.97 Phosphorus	12 S 32.06 Sulfur	13 Cl 35.45 Chlorine	14 Ar 39.95 Argon
5 Na 22.99 Sodium	6 Mg 24.31 Magnesium											15 Ga 69.72 Gallium	16 Ge 72.64 Germanium	17 As 74.92 Arsenic	18 Se 79.07 Selenium	19 Br 79.90 Bromine	20 Kr 83.80 Krypton
7 K 39.10 Potassium	8 Ca 40.08 Calcium											21 In 114.82 Indium	22 Sn 118.71 Tin	23 Sb 121.76 Antimony	24 Te 127.60 Tellurium	25 I 126.91 Iodine	26 Xe 131.29 Xenon
9 Rb 85.47 Rubidium	10 Sr 87.62 Strontium											27 Tl 204.38 Thallium	28 Pb 207.2 Lead	29 Bi 208.98 Bismuth	30 Po 209 Polonium	31 At 210 Astatine	32 Rn 222 Radon
11 Cs 132.91 Cesium	12 Ba 137.33 Barium											33 Hg 200.59 Mercury	34 Tl 204.38 Thallium	35 Pb 207.2 Lead	36 Bi 208.98 Bismuth	37 Po 209 Polonium	38 At 210 Astatine
13 Fr 223 Francium	14 Ra 226 Radium											39 Au 196.97 Gold	40 Pt 195.08 Platinum	41 Ag 107.87 Silver	42 Cd 112.41 Cadmium	43 In 114.82 Indium	44 Sn 118.71 Tin
												45 Rh 101.07 Rhodium	46 Pd 106.37 Palladium	47 Cu 63.55 Copper	48 Zn 65.38 Zinc	49 Ga 69.72 Gallium	50 Ge 72.64 Germanium
												51 Ir 223.03 Iridium	52 Pt 195.08 Platinum	53 Au 196.97 Gold	54 Hg 200.59 Mercury	55 Tl 204.38 Thallium	56 Pb 207.2 Lead
												57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids
												71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids

CLASSIFICATION

Metals	Nonmetals
Pre-transition	
Active	Corrosive
Alkali earth	
Alkaloids	
Transition	Intermediate
Post-transition	
Metals	
Metals (Ru-Pt, Os-Au)	Metals gas

57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids	57 to 70 Lanthanoids
71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids	71 to 102 Actinoids

Radioactive

Approximate
likely to be either
a metalloid or a
post-transition
metal

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals, those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalyst properties. Very weakly electropositive; inert to acids.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metals

Nonmetals

Metalloids

Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals

H and N are colorless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies; low to high electron affinities; moderate to high electronegativity values; relatively poor to moderately strong oxidising agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

*as black P, the most stable form in ambient conditions

Group 3

This starts as Sc-Y, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lf. Trends going down Sc-Y, and then

• La branch, resembles those of groups 1–2

• Lu branch, more resembles those of groups 4–10

Corrosive nonmetals

O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid, I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases

Colorless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answers

Arshad Iqbal



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