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## Student Exploration: Electron Configuration

### Activity B: Atomic radii

#### Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Na** (sodium).
- Select the ELECTRON CONFIGURATION tab.



**Question:** How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

**It will decrease**

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s2 2s2 2p6 3s1**

Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s2 2s2 2p6 3s2**

Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s2 2s2 2p6 3s2 3p1	118
Silicon	14	1s2 2s2 2p6 3s2 3p2	111
Phosphorus	15	1s2 2s2 2p6 3s2 3p3	98
Sulfur	16	1s2 2s2 2p6 3s2 3p4	88
Chlorine	17	1s2 2s2 2p6 3s2 3p5	79
Argon	18	1s2 2s2 2p6 3s2 3p6	71

# Periodic Table Exploration Answer Key

**D Keegan**



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## **Table of Contents Periodic Table Exploration Answer Key**

1. Understanding the eBook Periodic Table Exploration Answer Key
  - The Rise of Digital Reading Periodic Table Exploration Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Exploration Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Exploration Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Exploration Answer Key
  - Personalized Recommendations
  - Periodic Table Exploration Answer Key User Reviews and Ratings
  - Periodic Table Exploration Answer Key and Bestseller Lists
5. Accessing Periodic Table Exploration Answer Key Free and Paid eBooks
  - Periodic Table Exploration Answer Key Public Domain eBooks
  - Periodic Table Exploration Answer Key eBook Subscription Services
  - Periodic Table Exploration Answer Key Budget-Friendly Options
6. Navigating Periodic Table Exploration Answer Key eBook Formats

- ePub, PDF, MOBI, and More
  - Periodic Table Exploration Answer Key Compatibility with Devices
  - Periodic Table Exploration Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Periodic Table Exploration Answer Key
    - Highlighting and Note-Taking Periodic Table Exploration Answer Key
    - Interactive Elements Periodic Table Exploration Answer Key
  8. Staying Engaged with Periodic Table Exploration Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Periodic Table Exploration Answer Key
  9. Balancing eBooks and Physical Books Periodic Table Exploration Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Periodic Table Exploration Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Periodic Table Exploration Answer Key
    - Setting Reading Goals Periodic Table Exploration Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Periodic Table Exploration Answer Key
    - Fact-Checking eBook Content of Periodic Table Exploration Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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