

# Physical Science A

## Final Exam Study Guide

- **Matter**
  - Compare and contrast physical and chemical properties, and physical and chemical changes
  - Compare and contrast properties and behaviors of atoms in each of the 3 states of matter
  - Convert between mass, volume, and density, and relate appropriate SI units for each variable
  - Predict how solutions of different densities would settle in a graduated cylinder
- **Atoms, Elements, and the Periodic Table**
  - Describe Antoine Lavoisier and Dmitri Mendeleev's contributions to the development of the modern periodic table
  - Define the term ion and be able to predict the most common ions formed for elements by looking at the periodic table
  - Identify three subatomic particles and compare their properties
  - Compare and contrast atoms in ground and excited states
  - Distinguish the atomic number of an element from the mass number of an isotope, and use these numbers to describe the structure of atoms
  - Describe the currently accepted model of an atom
  - Interpret the information provided about the atoms and elements in a balanced chemical reaction
  - Differentiate between elements that are good and poor conductors of heat and electricity
  - Understand the development and organization of the periodic table, especially in terms of group properties, atomic mass and valence electrons
  - Identify general properties of metals, nonmetals, and metalloids
  - Define the words malleable, brittle, volatile, and ductile
  - Explain how the electron cloud model represents the behavior and locations of electrons in atoms
  - Predict the reactivity of some elements based on their locations in the periodic table
  - Compare and contrast Bohr and Rutherford's atomic models
  - Distinguish the atomic number of an element from the mass number of an isotope, and use these numbers to describe the structure of atoms
- **Bonding**
  - List the monomers that make up polymers for different biological macromolecules
  - Compare and contrast ionic and covalent bonds
  - Understand the development and organization of the periodic table, especially in terms of group properties, atomic mass and valence electrons
  - Explain the role that valence electrons play in bonding
  - Explain the octet rule for bonding
  - Draw Lewis dot diagrams for elements, and use them to predict bonding between one or more atoms
  - Identify three subatomic particles and compare their properties
  - Compare and contrast atoms in ground and excited states
- **Chemical Reactions**
  - Describe factors that can be used to increase reaction rates
  - Classify chemical reactions as synthesis, decomposition, single-replacement, double-replacement, or combustion reactions
  - Predict the products of chemical reactions
  - Interpret chemical equations in terms of reactants, products, and conservation of mass
  - Balance chemical equations by manipulating coefficients
  - Calculate amounts of reactants or products by using molar mass, mole ratios, and balanced chemical equations
  - Apply the Law of Conservation of Mass to chemical reactions
  - Use signs to tell that a chemical reaction has occurred
  - Distinguish between endothermic and exothermic reactions

# Note Taking Study Guide Physical Science Pages

**David Kirk**



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