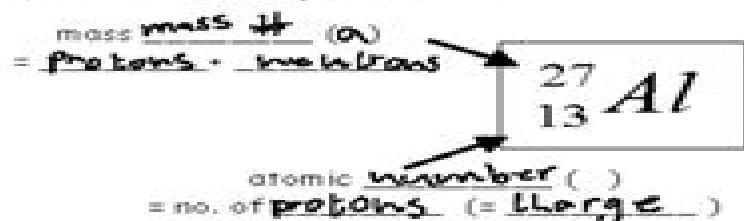


**Nucleons:**

- are particles occupying the nucleus
- consist of + charged protons and neutral charge neutrons
- have almost 2000 times the mass of electrons
- are made up of quark and leptons

**Nuclear Notation:**



${}_{13}^{27}\text{Al}$  has 13 protons and 14 neutrons, for a total of 27 nucleons.

${}_{13}^{28}\text{Al}$  has 13 protons and 15 neutrons, for a total of 28 nucleons.

**isotopes:** atoms of same element with different numbers of neutrons (different masses)

**Forces Acting on Nucleons:**

strong forces -

- forces of attraction between nucleons
- are independent of the charges of the nucleon
- are short range (exist only between closest neighbors)

electrical force -

- force of repulsion between + charged protons
- are long range

**When are nuclei unstable?**

- large nuclei ( $Z > 82$ ) - electrical forces of repulsion are greater than strong forces of attraction
- wrong neutron : proton ratio

**A radioactive isotope:**

- has an unstable nucleus
- spontaneously emits a particle and charge into another element

# Note Taking Guide Episode 4answers Chemistry

**Michael Brown**



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