

### Bohr's Energy Levels

- Electrons in \_\_\_\_\_
- \_\_\_\_\_ energy levels \_\_\_\_\_ to \_\_\_\_\_
- \_\_\_\_\_ energy levels \_\_\_\_\_ from \_\_\_\_\_
- Ground State: \_\_\_\_\_ in \_\_\_\_\_ possible

### Excited Atom

- Atom has \_\_\_\_\_
- \_\_\_\_\_ state is \_\_\_\_\_
- \_\_\_\_\_ soon \_\_\_\_\_ some amount of \_\_\_\_\_
- \_\_\_\_\_ seen as \_\_\_\_\_

### Wave Description of Light

Wavelength (\_\_\_\_\_): \_\_\_\_\_ between \_\_\_\_\_ on \_\_\_\_\_

Frequency (\_\_\_\_\_): the \_\_\_\_\_ of \_\_\_\_\_ waves \_\_\_\_\_ passing a given \_\_\_\_\_ in a given \_\_\_\_\_

$c = \frac{c}{\lambda \nu}$  : speed of \_\_\_\_\_

Sample problem #1:

What is the frequency of light if the wavelength is  $6.0 \times 10^{-7} \text{ m}$ ?

# Note Taking Guide Episode 303 Name Bohrs Energy Levels

**JL Elias**



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