

- 1) Arrange the following elements in order of increasing size: Al, B, C, K, and Na.

C	B	Al	Na	K
smallest				largest

- 2) Arrange the following elements in order of increasing size: Ca, Rb, P, Ge, and Sr.

P **Ge** **Ca** **Sr** **Rb**
smallest largest

- 3) Circle the atom or ion in each pair that has the larger radius.

- a) Cl or Cl
- b) Al or O
- c) In or I

- 4) Select the atom or ion in each pair that has the larger radius.

- a) Cs or Rb
b) O^{2-} or O
c) Br or As

- 5) Which of the following groups of elements is arranged correctly in order of increasing ionization energy?
(Answer by circling the correct answer.)

- a) $C < Si < Li < Ne$
 b) $Ne < Si < C < Li$
 c) $Li < Si < C < Ne$
 d) $Ne < C < Si < Li$

- 6) Arrange the following atoms in order of increasing ionization energy: Li, K, C, and N.

K Li C N
lowest highest

- 7) Arrange the following atoms in order of increasing ionization energy: Si, K, As, and Ca.

K	Cu	As	Si
lowest			highest

- 8) Compare the elements Na, Mg, O, and P.

- Which has the largest atomic radius? **Na**
- Which has the highest electron affinity? **O**
- Place the elements in order of increasing ionization energy.

Na	Mg	P	O
lowest			highest

Periodicity Worksheet Answer Key

Judd E. Hollander

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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