

Step 8. Periodic table trends

Electronegativity

Ionization energy

He

Atomic radius

The periodic table displays various elements with their symbols and atomic numbers. Trends are indicated by colored arrows:

- Electronegativity:** Indicated by a black arrow pointing right across the top.
- Ionization energy:** Indicated by a green arrow pointing right across the top and a green arrow pointing up on the right side.
- Atomic radius:** Indicated by a red arrow pointing down on the left side and a red arrow pointing left across the bottom.
- Non-metallic properties:** Indicated by a purple arrow pointing up and to the right.
- Metallic properties:** Indicated by a blue arrow pointing down and to the right.

Ionization Energy

Electronegativity

Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- Atomic radius of K is HIGHER than atomic radius of Br
- Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br



Periodic Trends Answer Key Section Review

Kaplan Test Prep



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