

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave better chemical properties of elements appeared in regular intervals and whereas elements were arranged according to increasing
 - (a) density
 - (b) atomic number
 - (c) atomic mass
 - (d) electronegativity
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of their atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The main reason the properties of the noble gases is that they are
 - (a) inert
 - (b) unreactive
 - (c) large
 - (d) far from reactivity
6. _____ Likewise, the first column in Group 1 has elements similar to H. These elements are the alkali metals, group 1a. The atomic number of
 - (a) 4
 - (b) 11
 - (c) 19
 - (d) 35
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 1a
 - (b) 2a
 - (c) 3a
 - (d) 4a
8. How many periods are in the table?
 - (a) 7
 - (b) 8
 - (c) 9
 - (d) 10
9. How many elements are in the group?
 - (a) 18
 - (b) 19
 - (c) 20
 - (d) 21
10. How many elements are in the group?
 - (a) 18
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11. How many elements are in the group?
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Minjie Lin

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