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Periodic Table Trends

Discovered Property	Definition
_____ 1. electron affinity	A. A measurement of the attraction of an atom for a pair of shared electrons of sharing with another atom.
_____ 2. electronegativity	B. A measurement of the amount of energy released when an electron is added to neutral atom.
_____ 3. atomic radius	C. The tendency of a neutral to react.
_____ 4. ionic radius	D. The tendency of a neutral to react.
_____ 5. ionization energy	E. A measurement of the size of a neutral atom.
_____ 6. metal reactivity	F. A measurement of the size of an ion.
_____ 7. metal reactivity	G. A measurement of the energy required to take away an electron from a neutral atom.

brown: electron affinity
 red: electronegativity
 blue: atomic radius
 purple: ionic radius
 orange: ionization energy
 green: numerical reactivity
 yellow: metal reactivity

lowest → highest

Periodic Table Trends Summary Answers Trends

**Dr Vijay Sarda, Dr A C Handa, Dr K K
Arora**



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Chemical Structure and Reactivity James Keeler, Peter Wothers, 2013-11 Why do certain substances react together in the way that they do What determines the shape of molecules And how can we predict whether a particular reaction will happen at all Such questions lie at the heart of chemistry the science of understanding the composition of substances their reactions and properties Though introductory chemistry is often broken into three sections inorganic organic and physical the only way for students to fully understand the subject is to see it as a single unified whole Chemical Structure and Reactivity rises to the challenge of depicting the reality of chemistry Offering a fresh approach to the subject by depicting it as a seamless discipline the text shows how organic inorganic and physical concepts can be blended together in order to achieve the common goal of understanding chemical systems With a lively and engaging writing style enhanced by vivid illustrations only Chemical Structure and Reactivity makes teaching chemistry with an integrated approach possible Special Features The only introductory text to take a truly integrated approach in explaining the fundamentals of chemistry Fosters an orbital based understanding of reactions with clear curly arrow mechanistic detail throughout A two part structure allows flexibility of use Part I lays down the core of the subject while Part II describes a series of relatively standalone topics which can be selected

to fit a particular course Numerous concepts are illustrated with fully cross referenced custom developed online modules enabling students to develop an understanding through active learning Self test exercises embedded in the text with solutions at the end of each chapter and extensive question sets encourage hands on learning to help students master the subject and gain confidence The Online Resource Centre features a range of additional resources for both students and registered adopters of the book New to this Edition A new chapter on symmetry has been added to Part I Discussions of organometallic chemistry spectroscopy and molecular geometry have been expanded Cross references from Part I to Part II have been increased to make the links between core concepts and more advanced topics clearer More self test questions and exercises have been provided

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Chemical Principles, Properties, and Reactions Kenneth L. Henold, Frank Walmsley,1984

Chemistry-vol-I Dr Vijay Sarda, Dr A C Handa, Dr K K Arora, A text book on Chemistry

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