



Percent Yield

45 grams of nitrogen monoxide (NO) react with an excess of sodium hydroxide (NaOH) following the balanced equation above. 6 grams of water were obtained from the reaction

What is the % yield?

Nitrogen monoxide (NO) gmm:

$$(1)(14.0) + (1)(16.0) = 30.0 \text{ g/mol}$$

$$\text{Actual Yield} = 6 \text{ g H}_2\text{O}$$

Water (H₂O) gmm:

$$(2)(1.0) + (1)(16.0) = 18.0 \text{ g/mol}$$

$$\text{Theoretical Yield} = 6.75 \text{ g H}_2\text{O}$$

$$\frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100 = 88.9\%$$

$$\frac{45 \text{ g NO} \times \frac{1 \text{ mol NO}}{30 \text{ g NO}} \times \frac{4 \text{ mol NO}}{4 \text{ mol NO}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \times 18 \text{ g H}_2\text{O}}{1} = 6.75 \text{ g H}_2\text{O}$$

Percent Yield Problems Answers

Padhraic Smyth

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Percent Yield Problems Answers:

Percent Yield Problems Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Percent Yield Problems Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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