Quantum Numbers

•

Used to ______ an _____ in an _____

42

Represents ______ energy level of ______

_____# of _____ in an

Example: What is the maximum number of electrons that can be in the _____ main energy level?

1

· This

Describes the ______ within an ______

of orbital _____possible in ______

Orbital Shapes

designated

- level 1: _____
- level 2; ______
- level 3: ______
- level 4: ______

How many electrons can each sublevel hold?

s = 1 orbital × 2 e /orbital = _____e

p = 3 orbitals x 2 e /orbital = _____e

d = 5 orbitals × 2 e /orbital = _____e

f = 7 orbitals x 2 e /orbital = ____e

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