

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

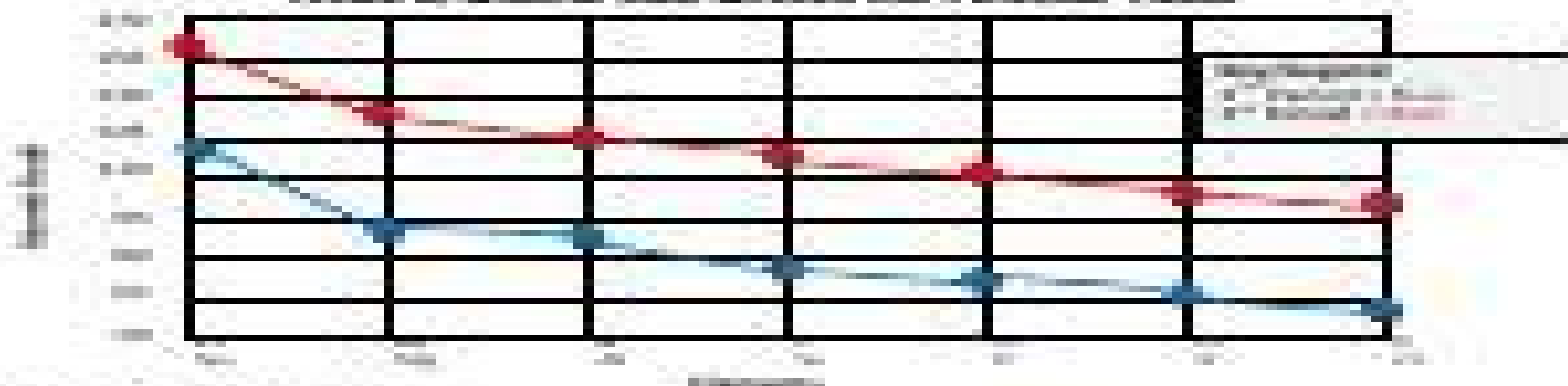
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, fill in the water graph.
3. Label the box as 'not suitable' (highlighted)
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
2	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
3	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
4	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
5	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
6	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
7	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius
8	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius	Electronegativity	Atomic radius

Now label the trends in the following table:

Trend in Atomic Size across the Periodic Table



Electronegativity Questions

1. Following the trend in atomic size across the periodic table, electronegativity
2. This means that as the atomic size increases, electronegativity decreases. This is because, as the atomic size increases, the distance between the nucleus and the outermost electrons increases.

Atomic Radius Questions

1. Electronegativity increases from left to right, i.e. from Li to F. This is because the distance between the nucleus and the outermost electrons decreases as the atomic size decreases.
2. This means that as the atomic size decreases, electronegativity increases. This is because, as the atomic size decreases, the distance between the nucleus and the outermost electrons decreases.

Periodicity Activity Okstate Answer Key

Wolfgang Guggemos

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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