

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Molar masses, stoichiometric ratios, limiting reagent, percent yield, theoretical yield, actual yield, percent error.

Chemists will often make reactions in which one reactant is in excess and the other is limiting. The limiting reagent is the one that runs out first.



The ratio required is 2:1 (hydrogen to oxygen) for the reaction to go.

2H ₂ (g)	20.0 g/mol
O ₂ (g)	32.0 g/mol
2H ₂ O(l)	36.0 g/mol
H ₂ O	18.0 g/mol

1. 20.0 g of each reactant will be used. How much H₂O will be produced from 20.0 g of each reactant? How many moles of hydrogen will be used? How many moles of oxygen will be used?



2. Which reactant is limiting? (The limiting reagent is the one that runs out first.)

3. How much water will be produced from 20.0 g of each reactant? (The limiting reagent will be used up first.)



4. If the reaction is performed in a lab, 22.5 g of water is produced from 20.0 g of each reactant. How much water is produced? (The limiting reagent will be used up first.)



Percent Yield And Limiting Reactants Quiz

Minjie Lin



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