

# THERMOCHEMICAL EQUATIONS AND TYPES OF ENTHALPY CHANGES



## Thermochemical Equations

A thermochemical equation is a reaction equation accompanied by an enthalpy change ( $\Delta H$ ). In addition to stating the number of moles reactants and the number of moles product, it also displays the amount of heat released or absorbed in the reaction in units of kJ.

### Example :

The thermochemical equation is known as:



This means that the formation of 1 mol  $\text{H}_2\text{O}$  hydrogen gas and oxygen gas released energy of 285,85 kJ (exothermic reaction).

Some things that must be considered in the thermochemical equation are

### 1. THE REACTION COEFFICIENT EXPRESSES MOLES.

#### Example:



(means that when 1 mol  $\text{H}_2$  reacts with 1 mol  $\text{I}_2$  to produce 2 mol  $\text{HI}$ , heat is absorbed by 51.9 kJ).

### 2. IF THE REACTION IS REVERSED, $\Delta H$ CHANGES SIGN

From reaction 1:



### 3. IF THE REACTIONS ARE MULTIPLIED, THEN $\Delta H$ IS MULTIPLIED

From reaction 1:



# Notetaking Guide For Thermochemical Equations

**SJ Ball**



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