

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- Boyle's Law

Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the collisions of the gas molecules impinging the surface.
- If the gas is compressible to small the volume it has large surface area and the pressure is very large.
- The large surface area is small on the walls of the container.
- Pressure of the gas will increase.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a pressure of 1.25 atm above the pressure inside the container. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 1.01 \text{ atm} \quad P_2 = 1.01 + 1.25 = 2.26 \text{ atm}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{800 \text{ ml} \times 1.01 \text{ atm}}{2.26 \text{ atm}} = 358 \text{ ml}$$

Ex. 2. A gas at pressure of 1.01 atm and 400 ml. of a certain gas has a volume of 350 ml. If the pressure is increased until the volume is 300 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 1.01 \text{ atm} \quad V_1 = 400 \text{ ml.} \quad V_2 = 300 \text{ ml.}$$

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{1.01 \text{ atm} \times 400 \text{ ml.}}{300 \text{ ml.}} = 1.35 \text{ atm}$$

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Scott C. Dulebohn



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