

Chemical Formula

Telling this problem involves and relates all atoms present in a compound.

Molecular $\text{C}_{10}\text{H}_{18}\text{O}$ (molecular)

Formula represents the number of atoms of each element in a single molecule.

Empirical $\text{C}_5\text{H}_9\text{O}$ (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound.

Atomic Mass

The average mass of all atoms of an element is weighted average.

Isotopes

^{12}C stable isotope carbon

^{13}C stable isotope carbon

Abundance

^{12}C abundance carbon

^{13}C abundance carbon

^{14}C abundance

Note Taking Guide Episode 6chemical Formula

R Sandford



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