

A. Identify an equilibrium, the chemical equation for that equilibrium, and its equilibrium constant expression.



Classroom activity: Ask students where the reaction is at equilibrium. Is it at equilibrium? What does that mean? If it is not at equilibrium, what does that mean? If it is at equilibrium, what does that mean?

B. Equilibrium:

• If the concentration of a reactant or product is not changing, the system is at equilibrium.

• If the concentration of a reactant or product is changing, the system is not at equilibrium.



$$K_{eq} = \frac{[C][D]}{[A][B]}$$

Eq. 1001: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.

$[A]$	0.100 M
$[B]$	0.100 M
$[C]$	0.100 M
$[D]$	0.100 M

Eq. 1001: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.



$$\frac{[H^+][OH^-]}{[H_2O]}$$

$$\frac{[H^+][OH^-]}{[H_2O]}$$

$$K_{eq} = 1.0 \times 10^{-14}$$

Note Taking Guide Episode 1202 Answers

K Morrison



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