

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 5.56 \text{ mol/kg} = 10.3^\circ C \quad \Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 5.56 \text{ mol/kg}$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of NaCl and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$. $\Delta T_f = 10.3^\circ C$, $\Delta T_b = 2.8^\circ C$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}} = 1.71 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 1.71 \text{ mol/kg} = 3.18^\circ C$$

3. How many grams of $C_6H_{12}O_6$ must be added to 100 g of H_2O to raise the boiling-point of the solution to $100.0^\circ C$? Assume that the solute does not dissociate. Equations: $K_b = 0.512^\circ C/m$

$$\Delta T_b = K_b \cdot m \quad 100.0^\circ C - 99.6^\circ C = 0.4^\circ C = 0.512^\circ C/m \cdot m$$

$$m = \frac{\Delta T_b}{K_b} = \frac{0.4^\circ C}{0.512^\circ C/m} = 0.781 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.781 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.0781 \text{ kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.781 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.0781 \text{ kg} = 78.1 \text{ g}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = 0.781 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = 0.781 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 100 g of water is prepared. The freezing-point of this solution is $-0.5^\circ C$. The normal freezing-point of pure water is $0^\circ C$ and the freezing-point depression constant for water is $1.86^\circ C/m$. Determine the molar mass of the compound. Assume it is a nonelectrolyte. Equations: $K_f = 1.86^\circ C/m$

$$\Delta T_f = K_f \cdot m \quad 0.5^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.269 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.269 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.0269 \text{ kg} = 26.9 \text{ g}$$

$$M = \frac{26.9 \text{ g}}{0.001 \text{ mol}} = 26.9 \text{ g/mol}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.269 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.0269 \text{ kg} = 26.9 \text{ g}$$

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$$\Delta T_f = K_f \cdot m \quad 0.5^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.269 \text{ mol/kg}$$

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