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4.14.

$$\left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.000000$$

$$\text{mass of H}_2\text{O} = 1.000000 \text{ g} \times \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.000000 \text{ g}$$

and that is the mass of the H<sub>2</sub>O molecule.

4.15.

- The water molecule has a mass of 18.01528 amu.
- The water molecule has a mass of 18.01528 amu.
- The water molecule has a mass of 18.01528 amu.

4.16.

From the mass of the H<sub>2</sub>O molecule, we find that the mass of the H<sub>2</sub>O molecule is 18.01528 amu.

$$\text{mass of H}_2\text{O} = 18.01528 \text{ amu} \times \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 18.01528 \text{ g}$$

4.17.

From the mass of the H<sub>2</sub>O molecule, we find that the mass of the H<sub>2</sub>O molecule is 18.01528 amu.

$$\text{mass of H}_2\text{O} = 18.01528 \text{ amu} \times \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 18.01528 \text{ g}$$

$$\text{mass of H}_2\text{O} = 18.01528 \text{ amu} \times \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 18.01528 \text{ g}$$

4.18.

From the mass of the H<sub>2</sub>O molecule, we find that the mass of the H<sub>2</sub>O molecule is 18.01528 amu.

$$\text{mass of H}_2\text{O} = 18.01528 \text{ amu} \times \left( \frac{1000 \text{ g}}{1000 \text{ g}} \right) = 18.01528 \text{ g}$$

# Physical Chemistry Ball Solution Manual

**Huangqi Zhang**



## Physical Chemistry Ball Solution Manual:

*Student Solutions Manual for Physical Chemistry* David Ball, 2002-09 The manual consists of complete solutions to all odd end of chapter exercises and problems

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