

Periodic Trends Worksheet

Directions: Use the trends in electronegativity, ionization energy, atomic radius, electron affinity, and metallic character to answer the following questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

K, Al, Na, S, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, aluminum, potassium.

O, S, Al, Na, K

3. Which atom has a higher first ionization energy than sodium?

None (sodium is the first ionization energy trend starts on the left)

4. Which two elements in the same family generally have similar properties?

There have been two cases here where a strong periodic trend is present

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which trend in the periodic table shows atomic size (or periodic table) to the greatest extent?

Electronegativity, ionization energy, atomic radius, and electronegativity

7. Which trend in ionization energy shows a significant and the periodic table? Which trend shows a significant trend?

Electronegativity, ionization energy, atomic radius, and electronegativity

8. Write the atomic radius of the following elements:

a. **Na** 186 pm
 b. **Al** 143 pm
 c. **S** 104 pm
 d. **O** 73 pm
 e. **K** 227 pm
 f. **Ca** 197 pm

a. **Na** 186 pm
 b. **Al** 143 pm
 c. **S** 104 pm
 d. **O** 73 pm
 e. **K** 227 pm
 f. **Ca** 197 pm

9. Write the atomic radius of the following elements:

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Lingsheng Yao

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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