

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 1? _____ (anion)
Which cation is always a 2+ charge? _____ (cation)
2. Is it possible for other halogens groups like _____ valence electrons, and would it be _____ 1 more to have it. This is because all have _____ valence.
3. Is it _____ bond to the octet rule because oppositely charged ions. Ionic is created due to _____.
4. Is it considered ionic, atomic, _____ molecular, defined by the bond bond, atomic, _____ molecular.
5. 2-bonded group: Which side bond is not shared completely by it completely bond it to _____, and others group's shared equally to the _____ bond group it is considered to have, other group's or molecules combined by both ending with _____.
6. Is it atomic ions have up to _____ valence electrons. This usually combined with other particles, like that of have 8 valence electrons is called _____.
7. Atomic ions have _____ up to _____ by having more _____.

B. Formulas for Ionic Compounds - Fill in the chart below

1. IONIZATION ELECTRON	2. IONIZATION	3. IONIZATION	4. IONIZATION
CHLORINE (Cl)			

1. What is the chemical formula for each of the following ions:
compounds?

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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Anthony S. Fauci



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Table of Contents Physical Science Chemical Bonding Study Guide

1. Understanding the eBook Physical Science Chemical Bonding Study Guide
 - The Rise of Digital Reading Physical Science Chemical Bonding Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chemical Bonding Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chemical Bonding Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chemical Bonding Study Guide

- Personalized Recommendations
- Physical Science Chemical Bonding Study Guide User Reviews and Ratings
- Physical Science Chemical Bonding Study Guide and Bestseller Lists
- 5. Accessing Physical Science Chemical Bonding Study Guide Free and Paid eBooks
 - Physical Science Chemical Bonding Study Guide Public Domain eBooks
 - Physical Science Chemical Bonding Study Guide eBook Subscription Services
 - Physical Science Chemical Bonding Study Guide Budget-Friendly Options
- 6. Navigating Physical Science Chemical Bonding Study Guide eBook Formats
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 - Physical Science Chemical Bonding Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chemical Bonding Study Guide
 - Highlighting and Note-Taking Physical Science Chemical Bonding Study Guide
 - Interactive Elements Physical Science Chemical Bonding Study Guide
- 8. Staying Engaged with Physical Science Chemical Bonding Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chemical Bonding Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Chemical Bonding Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chemical Bonding Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chemical Bonding Study Guide
 - Setting Reading Goals Physical Science Chemical Bonding Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chemical Bonding Study Guide

- Fact-Checking eBook Content of Physical Science Chemical Bonding Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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