

Average Atomic Mass

How can the masses of the particles within an atom be determined?

Objectives

After observing three gases, three ions, and naturally occurring three isotopes, the masses of these particles, ions, and isotopes should be determined. Comparison of particles including those belonging to the same element should show three different values resulting from isotopes of particles and elements having three different particles, neutrons. The relative mass of protons and neutrons will be made apparent by an experiment comparing naturally occurring isotopes of an element with hypothetical distributions.

Model 1 - A Strip of Magnesium Metal



1. Observe the distribution of the mass of Mg metal in Model 1.
How many ions does it contain? (Model 1)
2. Observe the mass distribution of the naturally occurring isotopes of magnesium (shown in Model 1) (A, B, C).
3. How all of the mass of magnesium in Model 1? Have the ions present mass? Explain.
Are there any mass values of the ions are different than mass values of the?
4. How the masses of the naturally occurring isotopes shown in Model 1? Show a table reflecting the mass distribution of the three isotopes and the masses of protons and neutrons present.

	^{24}Mg	^{25}Mg	^{26}Mg
Number of atoms	10	2	2

5. Calculate masses of magnesium in the mass distribution in Model 1.
 ^{24}Mg
6. Model in Model 1 show the ions are present in Model 1. How many of the mass of magnesium? Approximately how many naturally occurring ions are present? (Model 1)
As a group of 10 magnesium atoms, there will be approximately one ^{24}Mg , one ^{25}Mg , and eight ^{26}Mg atoms.

Periodicity Of Elements Pogil Key

Rachel Sandford



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