

# Light & Periodicity HOMEWORK

## Lesson 1: Homework

1. Match the correct term, listed to EACH symbol below:

- a.  $2.998 \times 10^8$   
 b. wavelength  
 c. frequency  
 d. energy  
 e.  $2.998 \times 10^{17}$

- a.  $6.626 \times 10^{-34}$   
 b. frequency  
 c. energy  
 d.  $2.998 \times 10^8$   
 e. Wavelength

2. Identify the units for EACH symbol below:

- a.  $c$  (m/s)  
 b.  $\lambda$  (m)  
 c.  $\nu$  (sec<sup>-1</sup>)

- a. meters (m)  
 b. 1/photon  
 c. sec<sup>-1</sup>  
 d. nanometers (nm)  
 e. 1/m  
 f. 1/m<sup>2</sup>

3. Convert  $6.65 \times 10^2$  nanometers to meters.

$$6.65 \times 10^2 \text{ nm} \times (1 \text{ m} / 10^9 \text{ nm}) = 6.65 \times 10^{-7} \text{ m}$$

4. As you orbit during orbit orbit around the sun, there is  $1.49 \times 10^8$  meters from Earth.

a) If a signal, which uses light waves, is sent that distance (a wavelength of  $1.49 \times 10^8$  meters), what is its frequency?

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8) &= (1.49 \times 10^8) \nu \\ \nu &= 2.01207986 \times 10^0 \text{ sec}^{-1} \end{aligned}$$

b) The units on the frequency is sec<sup>-1</sup>. If this value is inverted, the units are inverted as well. How many seconds does it take to send this signal? Give your answer with units from part "a" in this calculation.

$$(1 / 2.01207986 \text{ sec}^{-1}) = 0.000497014 \text{ sec} \approx 0.0005 \text{ sec}$$

5. Light can be used to excite the electrons in atoms. Certain frequencies of light work best for specific atoms. To excite an atom of Calcium, the ideal frequency of  $5.2 \times 10^{14} \text{ s}^{-1}$  is needed. What is the wavelength (in nm) of this electromagnetic radiation? Give your answer in scientific notation and with units.

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8) &= (5.2 \times 10^{14}) \lambda \\ \lambda &= 5.7653846 \times 10^{-7} \end{aligned}$$

$$5.7653846 \times 10^{-7} \text{ sec} \times 10^9 \text{ nm} = 576.53846 \text{ nm} \approx 576.54 \text{ nm}$$

6. The energy required to dissociate the  $\text{Cl}_2$  molecule into two separate Cl atoms requires a wavelength of  $5.00 \times 10^2$  meters. What is the energy of this light in joules?

$$\begin{aligned} 5.0000000 \times 10^2 \text{ meters} &= 500.00 \times 10^3 \text{ meters} \times 10^3 \text{ nm} = 500,000 \text{ nm} \\ &= 500,000 / 10^9 \text{ m} = 0.5 \text{ m} \end{aligned}$$

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8) &= (0.5 \times 10^3) \nu \\ \nu &= 5.996 \times 10^{14} \text{ sec}^{-1} \end{aligned}$$

$$\begin{aligned} E &= h \nu \\ (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) & \times (5.996 \times 10^{14} \text{ sec}^{-1}) \\ E &= 3.9730976 \times 10^{-19} \text{ Joules} \end{aligned}$$

7. Refrigerator coolant compounds (often compounds known as chlorofluorocarbons (CFCs), if not properly disposed, a refrigerator can leak this gas into the environment. Thus, ultraviolet radiation from the sun can destroy these CFCs, which can then destroy the ozone layer. How much energy do the CFCs receive from UV light that has a frequency of  $6.62 \times 10^{14} \text{ s}^{-1}$ . Give your answer in joules.

$$\begin{aligned} E &= h \nu \\ (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) & \times (6.62 \times 10^{14} \text{ sec}^{-1}) \\ E &= 4.386212 \times 10^{-19} \text{ Joules} \end{aligned}$$

8. You want your lawn sprinklers to come on in the morning, at first light. So, you design a switch that will operate by the photoelectric effect when enough sunlight hits it. As sunlight hits on your switch, the metal inside the device absorbs energy and ejects electrons to create electricity that powers the sprinklers. If the switch required  $5.8 \times 10^{14} \text{ s}^{-1}$  (just off metal to remove an electron), will the switch react if the sunlight has a wavelength of 540nm?

$$\begin{aligned} \nu &= c / \lambda \\ (2.998 \times 10^8) &= (540 \times 10^3) \nu \\ \nu &= 5.55185185 \times 10^{14} \text{ sec}^{-1} \end{aligned}$$

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8) &= (540 \times 10^3) \nu \\ \nu &= 5.55185185 \times 10^{14} \text{ sec}^{-1} \end{aligned}$$

$$\begin{aligned} E &= h \nu \\ (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) & \times (5.55185185 \times 10^{14} \text{ sec}^{-1}) \\ E &= 3.6800719 \times 10^{-19} \text{ Joules} \end{aligned}$$

$$\begin{aligned} (3.6800719 \times 10^{-19} \text{ Joules}) & > (3.6800719 \times 10^{-19} \text{ Joules}) \\ & > 3.68 \times 10^{-19} \text{ Joules} \end{aligned}$$

# Periodicity Ap Chemistry Answers

**Kathleen Armour**



## **Periodicity Ap Chemistry Answers:**

AP Chemistry Crash Course, Book + Online Adrian Dingle, 2025-04-03 AP Chemistry Crash Course A Higher Score in Less Time Make the most of your study time and earn a high score with America's bestselling rapid review for AP exams Here's why more AP students and teachers turn to REA's AP Chemistry Crash Course for today's AP digital exam Targeted Review Study Only What You Need to Know Our Crash Course is based on an in depth analysis of the latest AP Chemistry course and exam description and sample AP digital test questions released by the College Board It covers only the information tested on the exam keyed to the nine AP Chemistry units so you can make the most of your valuable study time Expert Test taking Strategies and Advice Written by Adrian Dingle an award winning AP Chemistry teacher and test development expert the book gives you the topics and critical context that will matter most on exam day Crash Course relies on the author's extensive analysis of the test's structure and content By following his advice you can boost your score Practice questions a mini test in the book a full length exam online Are you ready for your exam Try our focused practice set inside the book Then go online to take our full length practice exam You'll get the benefits of timed testing detailed answers and automatic scoring that pinpoints your performance based on the official AP exam topics so you'll be confident on test day Whether you're cramming for the exam or looking to recap and reinforce your teacher's lessons Crash Course is the rapid review study guide every AP student needs Avoid test anxiety Go with an author who knows exactly what you need to excel A true chemistry freelancer and subject matter expert author Adrian Dingle brings more than 30 years of classroom chemistry teaching experience along with tens of thousands of hours of one on one chemistry tutoring around the globe to a 17 year writing career that includes several bestselling international award winning chemistry books and a burgeoning portfolio of other chemistry writing projects

AP® Chemistry Crash Course, For the 2020 Exam, Book + Online Adrian Dingle, 2020-02-04 This Crash Course book is designed to be an aide for someone working to pass the AP advanced placement test in Chemistry

**CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

**AP® Chemistry Crash Course, 2nd Ed., Book + Online** Adrian Dingle, 2014-02-13 REA's Crash Course for the AP Chemistry Exam Gets You a Higher Advanced Placement Score in Less Time Crash Course is perfect for the time crunched student the last minute studier or anyone who wants a refresher on the subject Are you crunched for time Have you started studying for your Advanced Placement Chemistry exam yet How will you memorize everything you need to know before the test Do you wish there was a fast and easy way to study for the exam AND boost your score If this sounds like you don't panic REA's Crash Course for AP Chemistry is just what you need Our Crash Course gives you Targeted Focused Review Study Only What You Need to Know Fully revised for the 2014 AP Chemistry exam this Crash Course is based on an in depth

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