

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

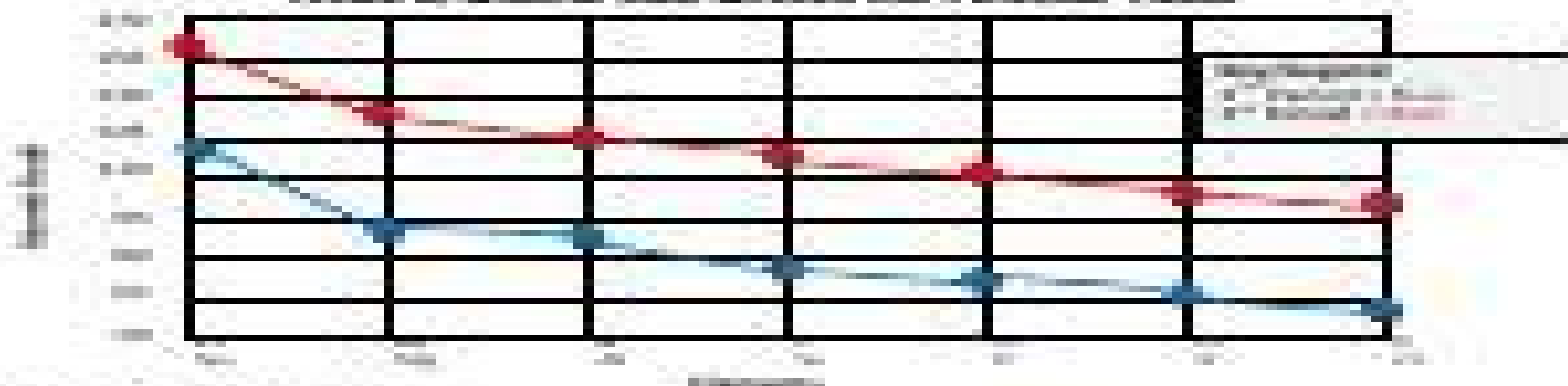
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, is it the other graph?
3. Label the box as 'and include ionization
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9Na	10Ne	11Mg	12Al	13Si	14P	15S	16Cl
3	17Ar	18Kr	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Ba	56La
8	57Ce	58Pr	59Nd	60Pm	61Sm	62Eu	63Gd	64Tb

Now label the elements above in the following way:

Trend in Atomic Size across the Periodic Table



Illustrated Trend Questions

1. Following the trend in atomic size across the periodic table decreases
2. This means two atoms that are more electronegative elements (fluorine) they are all smaller (smaller) meaning electronegativity related to ionization energy

Worked Trend Questions

1. Fluorine is more electronegative than oxygen (F is more electronegative than O) so it has a stronger attraction on electrons in atoms of the periodic table decreases
2. This means two atoms that are more electronegative elements (fluorine) they are all smaller (smaller) meaning electronegativity related to ionization energy

Periodicity Activity Okstate Answer Key

Michal Rosen-Zvi



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