

CHEMISTRY: PERIODIC LAW "SCRAMBLE" LAB

WHAT TO TURN IN:

Unscrambled table

Questions #1-6

Introduction

The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869. The basis of his arrangement was the atomic masses of the elements. This approach proved incorrect as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the basis of the atomic numbers of the elements. In accordance with Moseley's revision, the periodic law states the properties of the elements are periodic functions of their atomic numbers.

Each of the known elements has its own set of properties. These range from solid to gas, lustrous to dull, low to high melting points, various colors, etc. The elements are arranged within the periodic table into groups or families (vertical columns) and periods or rows (horizontal rows). This arrangement reflects the repeating nature of the properties of the elements.

In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table.

Objectives

- To arrange the elements in Groups IA-VIIIA (1, 2, 13-18) according to a list of clues and your knowledge of periodic properties
- To explain the trends of properties in groups and periods

Equipment: scissors, glue stick, blank table, scrambled table of elements

Procedure

1. Obtain a "scrambled" table. Each block on the table represents a different element from Groups IA-VIIIA, (1-18)
2. Cut out blocks A-Z. Use the clues given at the end of the procedure and arrange the elements in their proper order the blank table provided by your teacher. When you are sure of a correct placement, glue the square in place. Elements A-Z will fill the first four rows.
3. Cut out the remaining 16 blocks (these have a * instead of a code letter). Use the information provided in each block and your knowledge of periodic properties to arrange these elements in their proper position on the blank table. Glue the blocks in place in rows five and six. **THESE WILL BE AN EXTRA OPTIONAL GRADE AND WILL NOT AFFECT YOUR LAB SCORE IF YOU GET ANY INCORRECT OR DO NOT DO THEM.**
4. Some information is missing from each block. Predict the values for the missing items from the location of the element on the periodic table. Place your predictions on the table. (You may use a periodic table *only* to determine the symbol for each element.) You must include the element symbol, atomic number, and oxidation number (charge):

GROUP NUMBER

Group IA (1)
Group IIA (2)
Group IIIA (13)
Group IVA (14)
Group VA (15)
Group VIA (16)
Group VIIA (17)
Group VIIIA/0 (18)

GENERAL CHARGE (OXIDATION NUMBER) OF ION

1+
2+
3+
varies (mixed)
3-
2-
1-
0 (none)

Periodic Law Scramble Lab Answers

J Dewey



Periodic Law Scramble Lab Answers:

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Table of Contents **Periodic Law Scramble Lab Answers**

1. Understanding the eBook **Periodic Law Scramble Lab Answers**
 - The Rise of Digital Reading **Periodic Law Scramble Lab Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Periodic Law Scramble Lab Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Periodic Law Scramble Lab Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Periodic Law Scramble Lab Answers**
 - Personalized Recommendations
 - **Periodic Law Scramble Lab Answers** User Reviews and Ratings
 - **Periodic Law Scramble Lab Answers** and Bestseller Lists

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 - Periodic Law Scramble Lab Answers Public Domain eBooks
 - Periodic Law Scramble Lab Answers eBook Subscription Services
 - Periodic Law Scramble Lab Answers Budget-Friendly Options
6. Navigating Periodic Law Scramble Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Law Scramble Lab Answers Compatibility with Devices
 - Periodic Law Scramble Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Law Scramble Lab Answers
 - Highlighting and Note-Taking Periodic Law Scramble Lab Answers
 - Interactive Elements Periodic Law Scramble Lab Answers
8. Staying Engaged with Periodic Law Scramble Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Scramble Lab Answers
9. Balancing eBooks and Physical Books Periodic Law Scramble Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Scramble Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Law Scramble Lab Answers
 - Setting Reading Goals Periodic Law Scramble Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Law Scramble Lab Answers
 - Fact-Checking eBook Content of Periodic Law Scramble Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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