

MYP4 Chemistry

Periodic Trends

Use the periodic table, charts, and your knowledge of periodic trends to answer the following questions.

Atomic Radius

1. Which atom in each pair has the larger atomic radius? Circle your answer.

a) Li or **K** b) **Ca** or Ni c) **Ga** or B d) O or **C** e) Cl or **Br**
f) Be or **Ba** g) **Si** or S h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? From left to right in a period?

Atomic size (atomic radius) increases from top to bottom in a group and decreases from left to right in a period.

3. Why do atoms get smaller as you move left to right in a period?

Atoms get smaller as you move left to right in a period due to the increase in effective nuclear charge (while the number of occupied energy levels remain constant) leading to an increase in the force of attraction, which pulls the outer electrons closer to the nucleus and decreases the atomic radius.

4. Why do atoms get larger as you move down a group?

Atoms get larger as you move down a group due to the increase in occupied energy levels.

5. Arrange the following in order of increasing atomic radius atomic radius.

Mg, Na, P, Si and Ar? They are in the same period: Ar, P, Si, Mg, Na

Ionic Radius

6. Define an ion.

a charged atom or group of atoms

7. What is the difference between a cation and an anion?

Cations are positively charged ions. They are formed when an atom loses electrons.
Anions are negatively charged ions. They are formed when an atom gains electrons.

8. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

Cations are smaller than their neutral atom because they have less occupied energy levels.
Anions are larger than their neutral atom because of the increase in electron-electron repulsion.

9. Which particle has the larger radius in each atom/ion pair? Circle your answer.

a) **Na** Na⁺ b) S, **S²⁻** c) I, I⁻ d) **Al** Al³⁺

Periodic Trends Practice Problems Answer Key

Padhraic Smyth

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Periodic Trends Practice Problems Answer Key:

Problems and Problem Solving in Chemistry Education Georgios Tsaparlis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner **Prentice Hall Chemistry**, 2000 *E3 Chemistry Guided Study Book - 2018 Home Edition (Answer Key Included)* Effiong Eyo, 2017-12-08 Chemistry students and Homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Guided Study Book 2018 With E3 Chemistry Guided Study Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with guided step by step solutions to study and follow Practice multiple choice and short answer questions along side each concept to immediately test student understanding of the concept 12 topics of Regents question sets and 2 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN 978 1979088374 The Home Edition contains answer key to all questions in the book Teachers who want to recommend our Guided Study Book to their students should recommend the Home Edition Students and parents whose school is not using the Guided Study Book as instructional material as well as homeschoolers should also buy the Home edition The School Edition does not have the answer key in the book A separate answer key booklet is provided to teachers with a class order of the book Whether you are using the school or Home Edition our E3 Chemistry Guided Study Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year PLEASE NOTE Although reading contents in both the school and home editions are identical there are slight differences in question numbers choices and pages between the two editions Students whose school is using the

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