# Graphing Periodic Trends Lab

Name\_\_\_\_\_Hr\_\_\_

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

#### REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

ionization energy - The energy required to remove an electron from an atom.

#### DATA

Properties of the First 18 Elements

Properties of the First 18		Elements	
Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy
н	1	0.79	314
He	2	0.49	567
Li	3	2.05	124
Ве	4	1.4	215
8	5	1.17	191
C	6	0.91	260
M	7	0.75	335
0	8	0.65	314
F	9	0.57	402
No	10	7	2
Ma	1.1	2.23	119
Mg	12	1.72	176
A.I	13	1.82	138
Si	1.4	1.46	188
P	15	1.23	242
S	16	1.09	239
CI	17	0.97	299
Ar	18	?	2

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with guestions marks.

- Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)
- 2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)

## **Periodic Trends Lab Answers**

**LL Leslie** 

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