

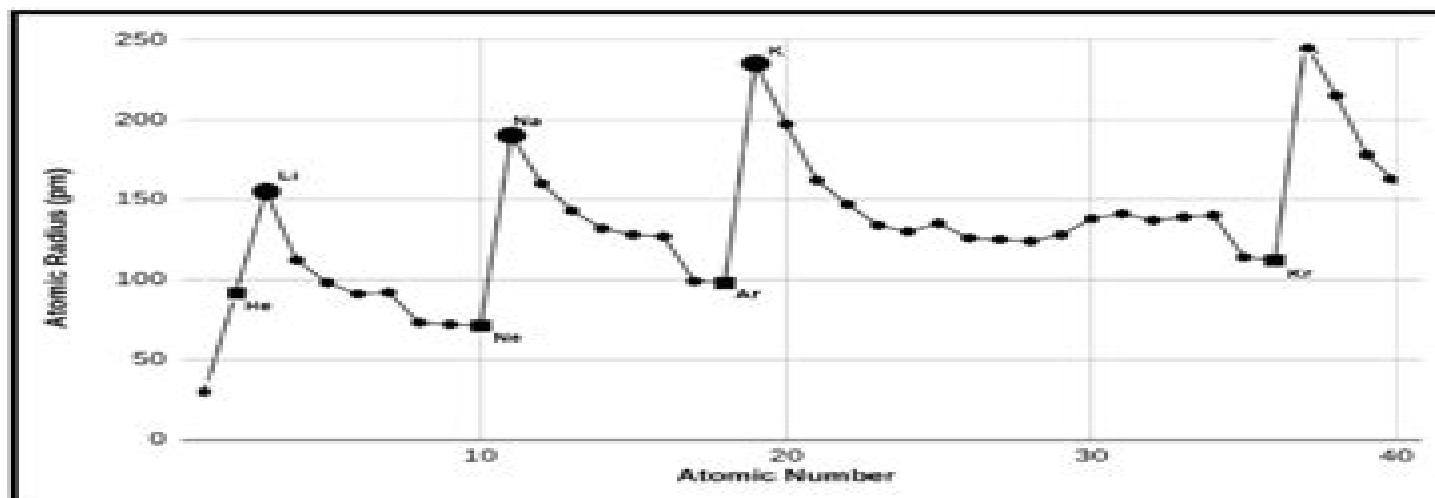


Figure 1.10: This graph shows the Atomic Radius (distance from the center of the nucleus to the valence (outer) level of electrons) vs Atomic Number, displaying the trend in atomic radius for the first 40 elements in the periodic table.

Description: Each successive period is displayed in a different section of the atomic number (x-axis), ex. 1-10, 11-20, 21-30, etc. As the atomic number increases within a period, the atomic radius decreases. To put simply, the peaks represent group 1 or alkali metals and the troughs represent group 18 or halogens.

DISCUSSION QUESTIONS: (Please answer on a separate document/paper)

PART ONE:

1. All the elements found at the peaks of your graph should fall into the same group.

a. What is the group number?

Group 1

b. What is the name of this group?

Alkali metals

2. All the elements found at the troughs of your graph should fall into the same group.

A. What is the group number?

Group 18

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Y Pai



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