

Name _____ Date _____ Per _____

Periodic Trends Lab

Periodic law states that when the elements are arranged in order of increasing atomic number, there is a periodic repetition of their chemical and physical properties. Elements within the same group have the most similar chemical and physical properties, although not identical. Elements in neighboring groups may also have similar chemical and physical properties, but to a lesser degree. Trends in properties occur within groups and across periods.

PURPOSE: In this activity, you will

- Interpret group and periodic trends in atomic radius, electronegativity, and ionization energy
- Graph these trends

PROCEDURE:

1. Complete the pre-lab section
2. Using the data from **Table S** of your reference tables, prepare the graphs listed below. Staple your graphs to the back of this lab packet.

On each graph be sure to:

- **USE PENCIL!**
- Put an appropriate title and labels for each axis.
- Omit the transition elements.
- Connect the points you plot.
- Connect the points representing Group 1 elements with a green line.
- Connect the points representing Group 2 elements with a blue line.
- Connect the points representing Group 17 elements with a red line.
- Draw a triangle around the points representing the elements of Period 3.

GRAPH 1. Atomic radius vs. atomic number for elements 1-54.

GRAPH 2. Electronegativity vs. atomic number for elements 1-54.

GRAPH 3. Ionization energy vs. atomic number for elements 1-54.

3. Complete the conclusion section.

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D Kirk



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Table of Contents Periodic Trend Lab

1. Understanding the eBook Periodic Trend Lab
 - The Rise of Digital Reading Periodic Trend Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trend Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trend Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trend Lab
 - Personalized Recommendations
 - Periodic Trend Lab User Reviews and Ratings
 - Periodic Trend Lab and Bestseller Lists
5. Accessing Periodic Trend Lab Free and Paid eBooks
 - Periodic Trend Lab Public Domain eBooks
 - Periodic Trend Lab eBook Subscription Services

- Periodic Trend Lab Budget-Friendly Options
- 6. Navigating Periodic Trend Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trend Lab Compatibility with Devices
 - Periodic Trend Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trend Lab
 - Highlighting and Note-Taking Periodic Trend Lab
 - Interactive Elements Periodic Trend Lab
- 8. Staying Engaged with Periodic Trend Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trend Lab
- 9. Balancing eBooks and Physical Books Periodic Trend Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trend Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trend Lab
 - Setting Reading Goals Periodic Trend Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trend Lab
 - Fact-Checking eBook Content of Periodic Trend Lab
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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