



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use permanent marker for writing and drawing.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol composed of one (A) or two (B) and list one of its intermolecular forces? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What (A) or (B) compound (A or B) has a (1-hexanol) greater heat of vaporisation? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 What compound (A or B) will have the higher surface tension? (2)

Compound B (M_r 102) is shown in the table. Use a suitable intermolecular force that both compounds A and B have to identify liquid boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound B is (1-hexanol) higher than compound A and B by referring to the ΔH_{vap} and intermolecular or intramolecular forces involved. (2)

3.2 Substances C1 is added to water whereas 1-pentanol (substance C2) are added to hexane. See also the diagram below.



- 3.2.1 Is substance C1 or C2, oil or water? (2)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the solubility in the intermolecular forces involved. (2)

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Eilon Morris



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