

## Physical Science Standards:

**SPS1:** Students will investigate our current understanding of the atom.

### a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
  - Proton has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Protons are also the elements \_\_\_\_\_ number of the identity of the element, so protons are found by looking at the atomic number
  - Neutron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Neutrons are found by subtracting the \_\_\_\_\_ Number from the \_\_\_\_\_ Mass or the BIG #- SMALL #.
  - Electron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Electrons are equal to the number of \_\_\_\_\_ in a neutral atom.
  - The nucleus is made up of these 2 particles, \_\_\_\_\_ and \_\_\_\_\_.
  - The electrons are located in this region of the atom, known as the \_\_\_\_\_.
  - The electron cloud is divided in to energy levels or energy shells, the first level will hold \_\_\_\_\_ electrons and the second level will hold \_\_\_\_\_ electrons.
  - The valence shell or valence level is the \_\_\_\_\_ energy shell of the atom.
- atomic mass and atomic number.
  - Atomic number is equal to the number of \_\_\_\_\_ in the atom and is also the atoms identity—it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
  - Define Isotope \_\_\_\_\_.
  - Ex. Carbon-12 and Carbon-14
    - Carbon 12 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
    - Carbon 14 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
- explain the relationship of the proton number to the element's identity.
  - The \_\_\_\_\_ number is equal to the number of \_\_\_\_\_ in the atom which is how the element is identified.
  - Ex. Nitrogen has an atom number of \_\_\_\_\_, which means that is has \_\_\_\_\_ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

# Physical Science Eoct Study Guide Forsyth

**JR Anderson**



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