

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

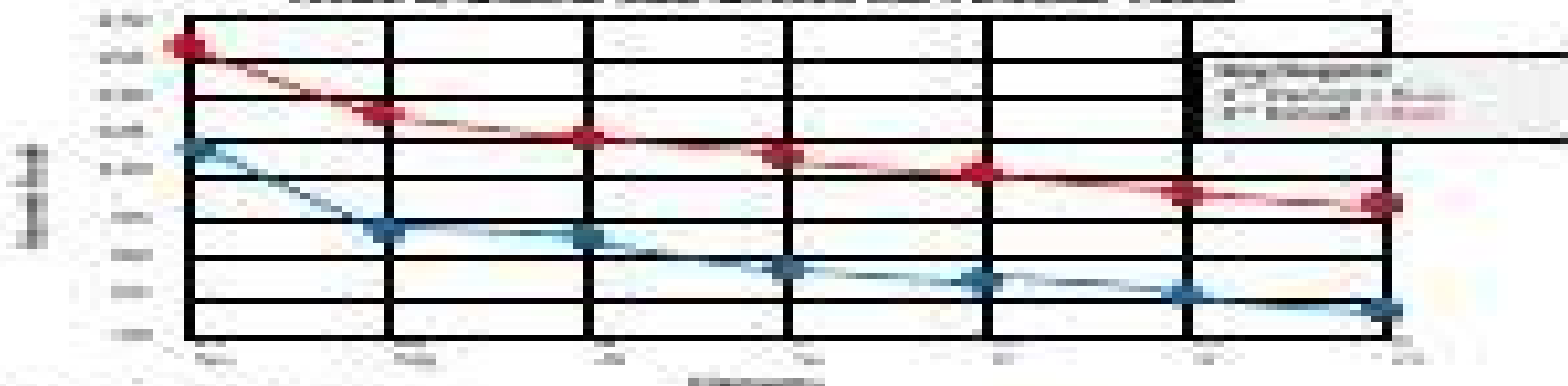
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, fill in the value graph
3. Label the line as acid reactivity increasing
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9Na	10Ne	11Mg	12Al	13Si	14P	15S	16Cl
3	17Ar	18Kr	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Ba	56La
8	57Ce	58Pr	59Nd	60Pm	61Sm	62Eu	63Gd	64Tb

Now label the trends above in the following way:

Trend in Atomic Size across the Periodic Table



Illustrated Trend Equations

1. Following the trend in atomic size across the periodic table decreases
2. This means that as the atomic size decreases, electronegativity increases (the opposite of electronegativity) meaning that the electronegativity increases

Worked Trend Equations

1. Electronegativity trend in the periodic table is decreasing (e.g. Na, K, Rb, Cs, Fr has the decreasing electronegativity as you go down in columns of the periodic table) decreasing
2. This means that as the atomic size increases, electronegativity decreases (the opposite of electronegativity) meaning that the electronegativity decreases

Periodicity Activity Okstate Answer Key

D Kirk



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Table of Contents Periodicity Activity Okstate Answer Key

1. Understanding the eBook Periodicity Activity Okstate Answer Key
 - The Rise of Digital Reading Periodicity Activity Okstate Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Activity Okstate Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodicity Activity Okstate Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Activity Okstate Answer Key
 - Personalized Recommendations
 - Periodicity Activity Okstate Answer Key User Reviews and Ratings
 - Periodicity Activity Okstate Answer Key and Bestseller Lists

5. Accessing Periodicity Activity Okstate Answer Key Free and Paid eBooks
 - Periodicity Activity Okstate Answer Key Public Domain eBooks
 - Periodicity Activity Okstate Answer Key eBook Subscription Services
 - Periodicity Activity Okstate Answer Key Budget-Friendly Options
6. Navigating Periodicity Activity Okstate Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodicity Activity Okstate Answer Key Compatibility with Devices
 - Periodicity Activity Okstate Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodicity Activity Okstate Answer Key
 - Highlighting and Note-Taking Periodicity Activity Okstate Answer Key
 - Interactive Elements Periodicity Activity Okstate Answer Key
8. Staying Engaged with Periodicity Activity Okstate Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Activity Okstate Answer Key
9. Balancing eBooks and Physical Books Periodicity Activity Okstate Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Activity Okstate Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodicity Activity Okstate Answer Key
 - Setting Reading Goals Periodicity Activity Okstate Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodicity Activity Okstate Answer Key
 - Fact-Checking eBook Content of Periodicity Activity Okstate Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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