

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Table Trend Activities Answer Key

RC Schank



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Table of Contents Periodic Table Trend Activities Answer Key

1. Understanding the eBook Periodic Table Trend Activities Answer Key
 - The Rise of Digital Reading Periodic Table Trend Activities Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Trend Activities Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Trend Activities Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Trend Activities Answer Key
 - Personalized Recommendations
 - Periodic Table Trend Activities Answer Key User Reviews and Ratings
 - Periodic Table Trend Activities Answer Key and Bestseller Lists

5. Accessing Periodic Table Trend Activities Answer Key Free and Paid eBooks
 - Periodic Table Trend Activities Answer Key Public Domain eBooks
 - Periodic Table Trend Activities Answer Key eBook Subscription Services
 - Periodic Table Trend Activities Answer Key Budget-Friendly Options
6. Navigating Periodic Table Trend Activities Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Trend Activities Answer Key Compatibility with Devices
 - Periodic Table Trend Activities Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Trend Activities Answer Key
 - Highlighting and Note-Taking Periodic Table Trend Activities Answer Key
 - Interactive Elements Periodic Table Trend Activities Answer Key
8. Staying Engaged with Periodic Table Trend Activities Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Trend Activities Answer Key
9. Balancing eBooks and Physical Books Periodic Table Trend Activities Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Trend Activities Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Trend Activities Answer Key
 - Setting Reading Goals Periodic Table Trend Activities Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Trend Activities Answer Key
 - Fact-Checking eBook Content of Periodic Table Trend Activities Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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