

Unit 2: The Periodic Table Review Sheet

U2L1: PT History

- On the present day Periodic Table, elements are arranged in order of increasing atomic number
- **Groups:** vertical columns on the periodic table
 - Elements in the same group have similar chemical properties - because they have the same number of valence electrons
 - Valence electron: outermost electrons (lost # in the electron configuration)
- **Periods:** horizontal rows on the periodic table

U2L2: Parts of the Periodic Table

- Elements are classified as either metals, nonmetals, or metalloids.
 - Classification can be determined by location on the periodic table or by different properties

Periodic Table of the Elements

Periodic Table of the Elements

1																	18				
1	H																	18			
Group																		Group			
1	Li	Be													13	B	C	N	O	F	Ne
2	Na	Mg													13	Al	Si	P	S	Cl	Ar
Group		3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	Group			
3	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr			
4	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe			
5	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn			
6	Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Uuq	Uup	Uuh	Uus	Uuo			

1

H

6

C

Atomic Mass → 12.011

Symbol → C

Atomic Number → 6

Electron Configuration → 1s² 2s² 2p²

Selected Oxidation States → -4, -3, -2, -1, 0, +1, +2, +3, +4

Bohr's Atomic Model and Quantum Numbers → 2, 4

Notes: Element is classified as a non-metal, located in the 2nd period of the periodic table.

Periodic Table Review Pt 2

General Electric Company

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the 'General Electric Company' text.

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