

Periodic Table Basics

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

Step 2: Determine the number of protons, neutrons, and electrons in each element.

Step 3: Create a Bohr diagram for each element.

Step 4: Draw the Lewis Structure for each element.

The card for Boron (B) contains the following information:

- Top left: Atomic number 5
- Center: Element symbol B and name Boron
- Bottom left: Atomic mass 10.81
- Top right: Protons (P) = 5, Neutrons (N) = 6, Electrons (E) = 5
- Bottom right: Bohr Diagram showing 2 shells with 2 electrons in the first shell and 3 electrons in the second shell.
- Bottom center: Lewis structure showing Boron with three valence electrons (represented by dots).

Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

Green = Li & Na

Pink = O & S

Blue = Be & Mg

Purple = F & Cl

Orange = B & Al

Red = C & Si

Tan = N & P

Yellow = He, Ne, & Ar

Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.



Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.

Periodic Table Basics Answer Key

Joacim Rocklöv



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