

Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: $[\text{Ne}] 3s^2 3p^2$ *may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an $s^2 p^3$ configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a $4s^2 4p^5$ electron configuration **Br, bromine**
- the 4th period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$ **O, oxygen**
- $4s^2$ **Ca, Calcium**
- $4s^2 3d^{10}$ **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

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The Princeton Review



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Table of Contents Periodic Table 6 Practice Problems Answer Key

1. Understanding the eBook Periodic Table 6 Practice Problems Answer Key
 - The Rise of Digital Reading Periodic Table 6 Practice Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table 6 Practice Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table 6 Practice Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table 6 Practice Problems Answer Key
 - Personalized Recommendations
 - Periodic Table 6 Practice Problems Answer Key User Reviews and Ratings
 - Periodic Table 6 Practice Problems Answer Key and Bestseller Lists

5. Accessing Periodic Table 6 Practice Problems Answer Key Free and Paid eBooks
 - Periodic Table 6 Practice Problems Answer Key Public Domain eBooks
 - Periodic Table 6 Practice Problems Answer Key eBook Subscription Services
 - Periodic Table 6 Practice Problems Answer Key Budget-Friendly Options
6. Navigating Periodic Table 6 Practice Problems Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table 6 Practice Problems Answer Key Compatibility with Devices
 - Periodic Table 6 Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table 6 Practice Problems Answer Key
 - Highlighting and Note-Taking Periodic Table 6 Practice Problems Answer Key
 - Interactive Elements Periodic Table 6 Practice Problems Answer Key
8. Staying Engaged with Periodic Table 6 Practice Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table 6 Practice Problems Answer Key
9. Balancing eBooks and Physical Books Periodic Table 6 Practice Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table 6 Practice Problems Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table 6 Practice Problems Answer Key
 - Setting Reading Goals Periodic Table 6 Practice Problems Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table 6 Practice Problems Answer Key
 - Fact-Checking eBook Content of Periodic Table 6 Practice Problems Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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