

POGIL: Periodic Trends in Atomic Properties

Information

Properties such as the size of an atom (atomic radius), the energy required to remove an electron from an atom (ionization energy), and the energy associated with the addition of an electron to a gaseous atom (electron affinity) can be understood in terms of the electron configuration of the atom and a competition between *electron-nucleus attraction* and *electron-electron repulsion*.

Preliminary Activity - Complete the orbital diagrams for the ground state configuration.

	1s	2s	2p		1s	2s	2p
H				C			
He				N			
Li				O			
Be				F			
B				Ne			

Model 1: Electron Ionization & Capture and Electrostatic Interactions

Reaction Equation for Electron Ionization: $X \rightarrow X^+ + e^-$ $\Delta E = 1^{\text{st}}$ ionization energy

Reaction Equation for Electron Capture: $X + e^- \rightarrow X^-$ $\Delta E = \text{electron affinity}$

Dominant inter-particle forces in helium:



Dominant inter-particle forces in lithium:



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U.S. Atomic Energy Commission



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, **Atomic Properties of the Elements**, 2001 *Understanding the Atom* U.S. Atomic Energy Commission, 1963
Chemistry and Our Universe, Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there

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Table of Contents Periodic Trends And Atomic Properties Pogil Answer Key

1. Understanding the eBook Periodic Trends And Atomic Properties Pogil Answer Key
 - The Rise of Digital Reading Periodic Trends And Atomic Properties Pogil Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends And Atomic Properties Pogil Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends And Atomic Properties Pogil Answer Key
 - User-Friendly Interface

4. Exploring eBook Recommendations from Periodic Trends And Atomic Properties Pogil Answer Key
 - Personalized Recommendations
 - Periodic Trends And Atomic Properties Pogil Answer Key User Reviews and Ratings
 - Periodic Trends And Atomic Properties Pogil Answer Key and Bestseller Lists
5. Accessing Periodic Trends And Atomic Properties Pogil Answer Key Free and Paid eBooks
 - Periodic Trends And Atomic Properties Pogil Answer Key Public Domain eBooks
 - Periodic Trends And Atomic Properties Pogil Answer Key eBook Subscription Services
 - Periodic Trends And Atomic Properties Pogil Answer Key Budget-Friendly Options
6. Navigating Periodic Trends And Atomic Properties Pogil Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends And Atomic Properties Pogil Answer Key Compatibility with Devices
 - Periodic Trends And Atomic Properties Pogil Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends And Atomic Properties Pogil Answer Key
 - Highlighting and Note-Taking Periodic Trends And Atomic Properties Pogil Answer Key
 - Interactive Elements Periodic Trends And Atomic Properties Pogil Answer Key
8. Staying Engaged with Periodic Trends And Atomic Properties Pogil Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends And Atomic Properties Pogil Answer Key
9. Balancing eBooks and Physical Books Periodic Trends And Atomic Properties Pogil Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends And Atomic Properties Pogil Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends And Atomic Properties Pogil Answer Key
 - Setting Reading Goals Periodic Trends And Atomic Properties Pogil Answer Key
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Periodic Trends And Atomic Properties Pogil Answer Key
 - Fact-Checking eBook Content of Periodic Trends And Atomic Properties Pogil Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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