

Chapter 5 The Periodic Table

Section 5.2 The Modern Periodic Table**(pages 130–138)**

This section explains the organization of the modern periodic table and discusses the general properties of metals, nonmetals, and metalloids.

Reading Strategy (page 130)

Previewing Before you read, complete the table by writing two questions about the periodic table on pages 132–133. As you read, write answers to your questions. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Questions About the Periodic Table	
Question	Answer
Students might ask: What does atomic mass mean?	Atomic mass is a value that depends on the distribution of an element's isotopes in nature and the masses of those isotopes.
Why are two columns of elements placed before the main body of the table?	Two kinds of elements are placed before the table to make the table more compact.

Other possible questions: Why are there two numbering systems for the columns? Why is period 7 incomplete?

The Periodic Law (pages 131–133)

1. Is the following sentence true or false? In the modern periodic table, elements are arranged by increasing number of protons.
_____ **true** _____
2. Explain why the number of elements per period varies. _____
_____ **The number of elements per period varies because the number of available orbitals increases from energy level to energy level.** _____
3. Properties of elements repeat in a predictable way when atomic numbers are used to arrange elements into groups. This pattern of repeating properties is called the _____
_____ **periodic law** _____

Atomic Mass (page 134)

4. Label the four types of information supplied for chlorine in the diagram.

a.	17
b.	Cl
c.	Chlorine
d.	35.453

- Atomic number**
- Element symbol**
- Element name**
- Atomic mass**

Periodic Table Section 5 2 Answer

Michael Brown

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Michael Brown.

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Table of Contents Periodic Table Section 5 2 Answer

1. Understanding the eBook Periodic Table Section 5 2 Answer
 - The Rise of Digital Reading Periodic Table Section 5 2 Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Section 5 2 Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Section 5 2 Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Section 5 2 Answer
 - Personalized Recommendations
 - Periodic Table Section 5 2 Answer User Reviews and Ratings
 - Periodic Table Section 5 2 Answer and Bestseller Lists
5. Accessing Periodic Table Section 5 2 Answer Free and Paid eBooks

- Periodic Table Section 5 2 Answer Public Domain eBooks
 - Periodic Table Section 5 2 Answer eBook Subscription Services
 - Periodic Table Section 5 2 Answer Budget-Friendly Options
6. Navigating Periodic Table Section 5 2 Answer eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Table Section 5 2 Answer Compatibility with Devices
 - Periodic Table Section 5 2 Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Table Section 5 2 Answer
 - Highlighting and Note-Taking Periodic Table Section 5 2 Answer
 - Interactive Elements Periodic Table Section 5 2 Answer
8. Staying Engaged with Periodic Table Section 5 2 Answer
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Section 5 2 Answer
9. Balancing eBooks and Physical Books Periodic Table Section 5 2 Answer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Section 5 2 Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Section 5 2 Answer
- Setting Reading Goals Periodic Table Section 5 2 Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Section 5 2 Answer
- Fact-Checking eBook Content of Periodic Table Section 5 2 Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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