

Name: _____

Atomic Structure & the Periodic Table Study Guide

Atomic Structure

Particle	Charge	Where are they found?
Proton	Positive	Nucleus
Neutron	Neutral/ no charge	Nucleus
Electron	negative	Electron cloud/energy levels outside nucleus

1. What is the rule for writing chemical symbols like Na or Ca? **Capitalize the first letter; lower case second letter**
2. What are the vertical columns on the periodic table called? **Groups** or **families**
 - a. How many of these are on the periodic table? **18**
 - b. Why are the elements in these columns? **They have the same chemical properties** (& same # of valence electrons)
 - c. Elements in the same group have the same number of **valence** electrons.
3. What are the horizontal rows on the periodic table called? **periods**
 - a. How many of these are on the periodic table? **7**
 - b. Elements in these rows have the same number of **energy** (electron shells) levels.
4. What are valence electrons? **Electrons in the outer energy level**
5. How many electrons are in each energy level?
1st energy level **2** 2nd energy level **8** 3rd energy level **8**
6. The atomic number is equal to the **# of protons** and also **# of electrons**
7. The atomic mass is equal to the **# protons + # of neutrons**
8. How do you find the number of neutrons? **Atomic mass – atomic number**
9. We view the order of the periodic table today according to the **atomic number**

Periodic Table Printable Study Guide

Scott C. Dulebohn



Periodic Table Printable Study Guide:

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Standards-based Activities and Assessments for the Differentiated Classroom Carolyn Coil,2004 How to plan and implement differentiation using practical strategies teacher friendly directions and time saving techniques

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