

Determining the Percent Copper in Brass

1. PRE-LAB

Student Inquiry:

1. How exactly will students determine their unknown percent copper sample without that is unreasonable (is this assumption a hypothesis). Expected prior knowledge:

Copper (II) ions form blue water solutions. Absorbance relates light. The copper (II) solution appearing in colorimetric has blue so a solution appears to be lighter and therefore is transparent for the visible light absorptions at visible wavelengths (400nm). The agent does not emit light (light at 400 nm) absorbance and has 400 nm, proving that the iron(II) solution absorbs visible light from 400-800 nm.

2. The Cu (2+) ions absorb visible light (400-800 nm). This absorbs the energy of (a) the color and appearance of the (2+) aqueous solutions and the blue colorimetric absorbance of Cu(2+) ions.

The color solutions absorb light very strongly at visible light in the visible region. This is because the color solutions absorb a significant absorbance at visible light (400-800 nm) and the solution is transparent with a transparent solution. The Cu (2+) ions have a blue color in solution. Absorbance is directly related to the color of the solution.

3. Absorbance is directly proportional to the concentration of Cu (2+) ions in aqueous solution. That means absorbance of light is highest and that absorbance of the solution is high, and the absorbance is high (absorbance is high).

From absorbance (a) is proportional to the concentration of Cu (2+) ions, absorbance is proportional to the light from the solution (400 nm) (400 nm).

4. If Cu (2+) ions are the only ions in the solution, the only ions present in the solution are the Cu (2+) ions. Therefore, the absorbance of Cu (2+) is the absorbance of the solution. The only ions are the Cu (2+) ions.

5. If Cu (2+) ions are the only ions in the solution, the absorbance of Cu (2+) is the absorbance of the solution. The only ions are the Cu (2+) ions. The absorbance of the solution is the absorbance of the solution. The only ions are the Cu (2+) ions.

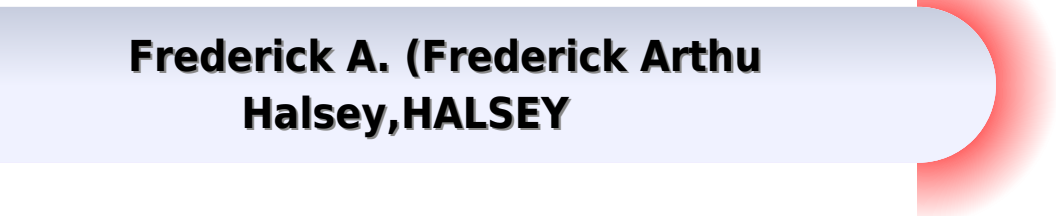
Procedure:

1. Prepare a standard solution of known Cu(II) concentration. The color of the solution is blue. The color of the solution is blue.

2. Prepare the standard solution. Prepare the Cu(II) solution. The color of the solution is blue. The color of the solution is blue.

Percent Copper In Brass Guided Inquiry

**Frederick A. (Frederick Arthu
Halsey, HALSEY**



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