



Class 10 Physical Science

Chemistry Formal Experiment: *Using flame colours to identify unknown metal ions*

August 2024

60 Marks

1 hour

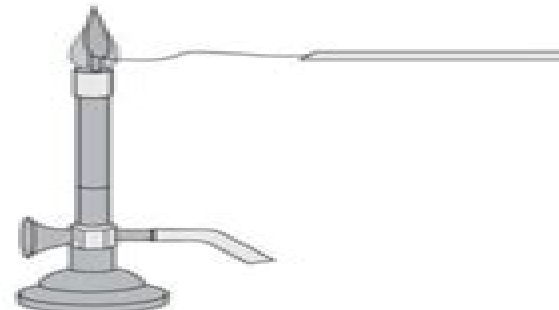
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Introduction:

The flame test is used to visually determine the identity of an unknown metal or metalloid, as an element or in a compound, based on the characteristic colour the salt turns the flame of a burner. When an electron in an excited state (higher energy) returns to its ground state (lower energy), an atomic spectrum is produced. A specific quantity of energy is associated with each electron, and when an electron returns to a lower energy level, a specific quantity of energy is released. The energy is released as electromagnetic radiation. Certain wavelengths will be visible to the naked eye; other wavelengths can only be detected with instruments. In the activity below, we will investigate the colours emitted by certain metal atoms. Different colours indicate different wavelengths of visible radiation.

Materials:

- ☐ NaCl(aq) , KCl(aq) , $\text{CaCl}_2 \cdot 2\text{H}_2\text{O(aq)}$, $\text{CuCl}_2 \cdot 2\text{H}_2\text{O (aq)}$,
- ☐ copper powder, magnesium powder,
zinc powder, iron powder
- ☐ four unknown salt solutions
- ☐ Bunsen burner
- ☐ Nichrome wire



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Raffaela Di Napoli



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Table of Contents Physical Science Formal Experiment Nomemo 2014

1. Understanding the eBook Physical Science Formal Experiment Nomemo 2014
 - The Rise of Digital Reading Physical Science Formal Experiment Nomemo 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Formal Experiment Nomemo 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Formal Experiment Nomemo 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Formal Experiment Nomemo 2014
 - Personalized Recommendations
 - Physical Science Formal Experiment Nomemo 2014 User Reviews and Ratings
 - Physical Science Formal Experiment Nomemo 2014 and Bestseller Lists
5. Accessing Physical Science Formal Experiment Nomemo 2014 Free and Paid eBooks

- Physical Science Formal Experiment Nomemo 2014 Public Domain eBooks
 - Physical Science Formal Experiment Nomemo 2014 eBook Subscription Services
 - Physical Science Formal Experiment Nomemo 2014 Budget-Friendly Options
6. Navigating Physical Science Formal Experiment Nomemo 2014 eBook Formats
- ePub, PDF, MOBI, and More
 - Physical Science Formal Experiment Nomemo 2014 Compatibility with Devices
 - Physical Science Formal Experiment Nomemo 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Formal Experiment Nomemo 2014
 - Highlighting and Note-Taking Physical Science Formal Experiment Nomemo 2014
 - Interactive Elements Physical Science Formal Experiment Nomemo 2014
8. Staying Engaged with Physical Science Formal Experiment Nomemo 2014
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Formal Experiment Nomemo 2014
9. Balancing eBooks and Physical Books Physical Science Formal Experiment Nomemo 2014
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Formal Experiment Nomemo 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Formal Experiment Nomemo 2014
- Setting Reading Goals Physical Science Formal Experiment Nomemo 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Formal Experiment Nomemo 2014
- Fact-Checking eBook Content of Physical Science Formal Experiment Nomemo 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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