Chapter 5 Study Guide- Physical Science Test Monday, April 28th, 2014

- What is a chemical bond? force of attraction that holds elements together in order to form a unit
- Define chemical structure, the arrangement of atoms in a substance.
- What is an ionic bond? a chemical bond in which one atom loses an electron to form a positive ion and the other atom gains an electron to form a negative ion
- 4. An ionic bond is between a <u>METAL</u> and a <u>NONMETAL</u>.
- What is a covalent bond? a chemical bond that involves sharing a pair of electrons between atoms in a molecule
- 6. A covalent bond is between a nonmetal and a nonmetal .
- What is a metallic bond? a bond formed by the attraction between positively charged metal ions and the electrons around them.
- 8. What are the polyatomic ions? an ion made of two or more atoms.
- What is an ION? a charged atom
- 10. What is a CATION? A positively charged atom
- 11. What is an ANION? A negatively charged atom
- 12. Which part of an ionic compound should be written first? Metal

Name the following ionic compounds:

- 13. Fel2- Iron Iodide
- 14. MnF3- Manganese Fluoride
- 15. CrCl2- Chromium Chloride
- 16. CuS- Copper Sulfide

Name the following covalent compounds:

- 17. As2Os- diarsenic pentaoxide
- 18. Sils- silicon tetrafodide
- 19. P4S3- tertraphosphorus trisulfide
- 20. PaOin- tetraphosphorus decoxide

Write the equation for the following ionic compounds:

- 21. Lithium Oxide- LizO
- 22. Magnesium Nitride- Mg3N2
- Sodium Phosphide- NasP
- 24. Calcium Bromide CaBrz

Define/Explain the following

- 25. What is a subscript? number written below, tells you how many of each element is present in compound/molecule
- 26. What is a superscript? number written above the element symbol, this is how the oxidation number (charge) is displayed.

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